

## Experiment: Charles's Law

### Introduction:

According to Charles' Law the volume of a gas is directly proportional to the Kelvin temperature assuming the pressure is constant or

$$V = kT$$

Where,  $V$  = volume,  $T$  = the Kelvin temperature and  $k$  = a constant number

If we have two different temperatures,  $T_1$  and  $T_2$  Charles Law can be restated as:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Using Charles Law you will predict the volume the gas should be at the lower temperature based on its volume at the higher temperature. Then you will compare the predicted volume at the lower temperature with the actual measured volume.

To summarize, a given volume of air was taken at the temperature of boiling water. The air was cooled, causing it to contract. The new volume was determined experimentally.

### Objective:

The purpose of this experiment is to study the changes in the volume of a gas with changes in temperature at constant pressure.

### Apparatus:

Erlenmeyer Flask, 250 mL, Thermometer, Beaker, 600 mL, Graduated cylinder 100 mL, Rubber stopper fitted with a pinch clamp, Large dish pan, Ring stand and clamp, Hot plate.

### Procedure (labelled flow diagram):

# Lab Manual Charles Law

**J. A. Beran, Mark Lassiter**



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