

Procedure

You will use the simulation on the following website to virtually measure and record all voltages for this experiment.

<http://web.mst.edu/~gbert/Electro/Electrochem.html>

Once you have gone to the website above you should see the interface below.



- You will be able to adjust the electrodes and solutions using the drop down menus on the right and left. The simulator will only allow you to use a solution of the metal salt corresponding to the metal electrode placed in the beaker. (For example: Copper(II) nitrate must be used with the copper electrode.)
- You can also adjust the concentration of each of the metal salt solutions in the beakers by typing the value in the box on either side of the simulator.
- To measure the voltage of any cell that you have constructed, simply click **Measure Cell Voltage** on the bottom left of the simulator. If there is a problem with the cell, the simulator will identify any errors in your construction.
- You will not be using the **New Problem** link in the or "Level" drop down menu in the lower right for this assignment.

Part A

1. You will be measuring the cell potentials for all the possible combinations of each of the following electrodes: Ag/Ag^+ , Cu/Cu^{2+} , Ni/Ni^{2+} , Pb/Pb^{2+} , and Zn/Zn^{2+} . Construct the cells so that the metal salt solution in each beaker has a concentration of 0.10 M and a positive voltage is read on the voltmeter when the cell potential is measured.
2. Place a silver electrode in the beaker on the left using the top drop down menu and then select silver nitrate in the second drop down menu and type in a concentration of 0.10 into the box at the bottom.

Lab Manual For Chem 1212l

JE Gale



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Table of Contents Lab Manual For Chem 1212l

1. Understanding the eBook Lab Manual For Chem 1212l
 - The Rise of Digital Reading Lab Manual For Chem 1212l
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Manual For Chem 1212l
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Manual For Chem 1212l
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Manual For Chem 1212l
 - Personalized Recommendations
 - Lab Manual For Chem 1212l User Reviews and Ratings
 - Lab Manual For Chem 1212l and Bestseller Lists

5. Accessing Lab Manual For Chem 1212l Free and Paid eBooks
 - Lab Manual For Chem 1212l Public Domain eBooks
 - Lab Manual For Chem 1212l eBook Subscription Services
 - Lab Manual For Chem 1212l Budget-Friendly Options
6. Navigating Lab Manual For Chem 1212l eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Manual For Chem 1212l Compatibility with Devices
 - Lab Manual For Chem 1212l Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Manual For Chem 1212l
 - Highlighting and Note-Taking Lab Manual For Chem 1212l
 - Interactive Elements Lab Manual For Chem 1212l
8. Staying Engaged with Lab Manual For Chem 1212l
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Manual For Chem 1212l
9. Balancing eBooks and Physical Books Lab Manual For Chem 1212l
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Manual For Chem 1212l
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lab Manual For Chem 1212l
 - Setting Reading Goals Lab Manual For Chem 1212l
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab Manual For Chem 1212l
 - Fact-Checking eBook Content of Lab Manual For Chem 1212l
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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