

MCAT Chemistry Equation Sheet

ORBITALS AND ELECTRONS

Quantum numbers descriptions

Quantum mamber	Description	Possible values
Principal quantum manber (n)	energy level	n=1, 2, 3
Angular monocotum quantum number (f)	orbital shape (s, p, d, f)	J = 0, sos- 1
Magnetic quantum number (m)	orbital oriomation (p., p., p.)	and the state of
Spile quantum number (m.)	describes spin of electron	100, 11 11 115 OK

Each electron has a unique set of quantum numbers.

ELECTRONIC CONFIGURATION

Electron configuration: describes the location of all of the electrons in an element, e.g. Carbon = ta/2a/2bc/

Heisensberg Unicertainty Principle: the position and monopolism of a particle cannot booth be exactly known at the same time. Pauli exclusion principle: no two electrons in an attein can have the same set of floor quantum numbers (n, L, m, m,) which means that each orbital with quantum numbers (n, L, m) can hold at most two electrons with exposition spins.

Hand's rule: when electrons added to orbitals of equal energy, a single electron enters each orbital before a second electron with appealie spin enters any orbital. The spins romain parallel if possible, e.g.

Notrogem N = T	74	7.4	1.1
T.	24	24	2p., 2p.,

Audhen Principle: dictates the order in which orbitals fill. Some exceptions occur in the regunition elements.

1s 2s 2p 3s 3p 3d 4s 4p 4d 4f 5s 5p 5d 5f 6s 6p 6d 6f

THE GAS PHASE

Units of Pressure:

Table = 760 rhmHg = 760 torr = 101.325 kPa = 1.01325 barr

CONTRACTORS WORTHWATERS

- molecules are very small compared to the distance b'w them.
- 2. molocules are continuously moving
- pressure of the gas collisions of the meteories with container walls
- 4. molecules do not experience intermolecular

Ideal-gas law: PV = nRT

Combined gas less: $\frac{P_i V_c}{T_c} \approx \frac{P_i V_j}{T_c}$

Charles' law: $\frac{V_i}{T_i} = \frac{V_i}{T_i}$

 $T_i = T_i$ Heatic's how: $BK = B_i V_i$ T = qq

No ones Law $P_1 = P_2$ V = cons

STP – standard temperature (0°C or 273.15 K) and movement (1 stan)

I mol of ideal gas occupies F = 22.4 L standard conditions - 25°C and 1 atm

Distance's law of partial pressure (mixture of the gases)

$$P = p_A + p_B + ... + p_n$$

 $p_A = \gamma_A \cdot P$

$$\chi_x = \frac{n_x}{n_{\text{max}}} = \frac{N_x}{N_{\text{max}}}$$

GRAHAM'S LAW OF EFFUSION

$$\frac{v_s}{v_s} = \sqrt{\frac{m_s}{m_s}}$$

$$KE_A = \frac{1}{2}m_Av_A^2$$

$$KE_A = KE_B = T_A = T_B$$
 bic $KE \propto T$, but

$$v_A \neq v_B \neq T_A = T_B$$

What leads to the pas "idealty", high T & low P & lower molecular weight or small size of gas molecules

Real cases:

- have intermolecular attractions (non-elastic optisions) called Van der Waals forces.
- contain molecules that have volume (or make make site)

$$\left(P_{nm} + a \frac{n^2}{V^2}\right) \cdot \left(V_{nm} - nb\right) = nRT$$

INTERMOLECULAR FORCES

DISPERSION FORCES:

Describe the attraction between the negatively charged electron cloud of one molecule and the positively charged nuclei of neighbouring molecules Depos Ar Forces:

Describe the attraction between the negatively charged end of a polar molecule and the positively charged ends of neighbouring polar molecules Hydeopen Bosones:

Involves lone pairs of electrons on an electronegative atom of one molecule and a polar bond to hydrogen in another molecule, they are confined to molecules that contain O, N, and F stores.

Procamicanicity:

The ease with which the electron clouds are distorted

MOLEDULAR SOLIES

Molecules in molecular solids are held in place by the types of forces: dispersion forces, dipolar interactions, and/or hydrogen bonds

METALLIC SOLICE

Atoms in metallic solids are held in place by

delocalised bonding

Nameons Sould:

Contains an array of covalent bonds linking every atom to its neighbours

JONES SOLID:

Contains cations and anions, attracted to one another by coulombic interactions

SOLUTIONS

molarity (M)	modes of solute/volume of solution. M=s/V
modulating days	moles of solute/kg of solvent
mode fraction (g)	medies of solute/total motes
mass percent	mass of solute x 190/ total mass of solution
parts per million (ppm)	mass of solute s 10°/total mass of solution

Reputt's Law (distillation)

$$P_x = X_x P_x^{\alpha}$$

Where P_d = partial pressure of A in solution X_d = mode fraction of A in the liquid P_d^* = vapour pressure of pure A

Colligative properties (melting/boiling points etc.)

Frocering point depression; salt on the sidewalk in winter lowers the focuring point of H₂O to prevent the water from froceing

$$\Delta T_{\nu} = 3c_{\nu} m_{\nu}$$

Booling point elevation: salt in pasts water increases the booling point of H_O so that the water bools hotter (and year pasts cooks more queckly).

$$\Delta T_{\rm in} = k_{\rm inversion}$$

Outnotic pressure: increase in pressure due to a solvent crossing a membrate into a restry concentrated solution. IIV—edit7



Mcat Chemistry Review Sheet

Alexander Stone Macnow

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