



Redox Reactions

Name _____
Date _____

Balance the following redox reactions.

1. $\text{H}_2\text{SO}_3 + \text{I}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4 + \text{HI}$
2. $\text{FeCl}_3 + \text{SnCl}_2 \rightarrow \text{FeCl}_2 + \text{SnCl}_4$
3. $\text{SbCl}_5 + \text{KI} \rightarrow \text{SbCl}_3 + \text{KCl} + \text{I}_2$
4. $\text{CdS} + \text{I}_2 + \text{HCl} \rightarrow \text{CdCl}_2 + \text{HI} + \text{S}$
5. $\text{KClO}_3 + \text{FeSO}_4 + \text{H}_2\text{SO}_4 \rightarrow \text{KCl} + \text{Fe}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
6. $\text{CuSO}_4 + \text{KI} \rightarrow \text{CuI} + \text{K}_2\text{SO}_4 + \text{I}_2$
7. $\text{AuCl}_3 + \text{KI} \rightarrow \text{AuCl} + \text{KCl} + \text{I}_2$
8. $\text{SnCl}_2 + \text{HgCl}_2 \rightarrow \text{SnCl}_4 + \text{HgCl}$
9. $\text{Ti}_2(\text{SO}_4)_3 + \text{FeI}_2(\text{SO}_4)_3 \rightarrow \text{Ti}(\text{SO}_4)_2 + \text{FeSO}_4$
10. $\text{NaClO} + \text{H}_2\text{S} \rightarrow \text{NaCl} + \text{H}_2\text{SO}_4$
11. $\text{Fe}_3\text{O}_4 + \text{KMnO}_4 + \text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O}$
12. $\text{NaCrO}_2 + \text{NaOH} + \text{H}_2\text{O}_2 \rightarrow \text{Na}_2\text{CrO}_4 + \text{H}_2\text{O}$
13. $\text{Ag} + \text{HNO}_3 \rightarrow \text{AgNO}_3 + \text{NO} + \text{H}_2\text{O}$
14. $\text{Ag}_2\text{S} + \text{HNO}_3 \rightarrow \text{AgNO}_3 + \text{NO} + \text{S} + \text{H}_2\text{O}$

(continued)



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Judd E. Hollander



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Table of Contents J Weston Walch Publisher Crossword

1. Understanding the eBook J Weston Walch Publisher Crossword
 - The Rise of Digital Reading J Weston Walch Publisher Crossword
 - Advantages of eBooks Over Traditional Books
2. Identifying J Weston Walch Publisher Crossword
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an J Weston Walch Publisher Crossword
 - User-Friendly Interface
4. Exploring eBook Recommendations from J Weston Walch Publisher Crossword
 - Personalized Recommendations
 - J Weston Walch Publisher Crossword User Reviews and Ratings

- J Weston Walch Publisher Crossword and Bestseller Lists
- 5. Accessing J Weston Walch Publisher Crossword Free and Paid eBooks
 - J Weston Walch Publisher Crossword Public Domain eBooks
 - J Weston Walch Publisher Crossword eBook Subscription Services
 - J Weston Walch Publisher Crossword Budget-Friendly Options
- 6. Navigating J Weston Walch Publisher Crossword eBook Formats
 - ePub, PDF, MOBI, and More
 - J Weston Walch Publisher Crossword Compatibility with Devices
 - J Weston Walch Publisher Crossword Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of J Weston Walch Publisher Crossword
 - Highlighting and Note-Taking J Weston Walch Publisher Crossword
 - Interactive Elements J Weston Walch Publisher Crossword
- 8. Staying Engaged with J Weston Walch Publisher Crossword
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers J Weston Walch Publisher Crossword
- 9. Balancing eBooks and Physical Books J Weston Walch Publisher Crossword
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection J Weston Walch Publisher Crossword
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine J Weston Walch Publisher Crossword
 - Setting Reading Goals J Weston Walch Publisher Crossword
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of J Weston Walch Publisher Crossword
 - Fact-Checking eBook Content of J Weston Walch Publisher Crossword
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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