

**Finishing Part A**  
Use the chart below to determine the pH of each solution. Each pH indicator changes color based on the pH of the solution. Look at Methyl Red as an example. Methyl red will be red at and below a pH of 4.8. Additionally, methyl red will be yellow at or above a pH of 6.0. In between a pH of 4.8 and 6.0, methyl red will be orange (red and yellow make orange).

Each pH indicator has a limited range. Use the combination of the indicators to narrow down the pH range for each solution. For example, consider a solution with unknown pH. Start with the indicators that change color around a pH of 7.0, which will be bromothymol blue and phenol red. Our bromothymol blue indicator is blue, so the pH must be greater than 7.6. Likewise, the phenol red indicator is red, so the pH must be greater than 8.0. However, our phenolphthalein solution is colorless, indicating a pH lower than 8.2. Therefore, we have determined that the solution must have a pH between 8.0 and 8.2.

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1. Record the experimental pH range for each solution below

Distilled Water \_\_\_\_\_  
Tap Water \_\_\_\_\_  
0.1M NaOH \_\_\_\_\_  
0.1M  $\text{NH}_4\text{OH}$  \_\_\_\_\_  
0.1M HAc \_\_\_\_\_  
0.1M HCl \_\_\_\_\_  
0.1M  $\text{H}_2\text{SO}_4$  \_\_\_\_\_

2. Calculate the theoretical pH for each solution and record it below.
- Distilled Water

# Gsu Chemistry 1212 Lab Manual

**Scott C. Dulebohn**



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## **Table of Contents Gsu Chemistry 1212 Lab Manual**

1. Understanding the eBook Gsu Chemistry 1212 Lab Manual
  - The Rise of Digital Reading Gsu Chemistry 1212 Lab Manual
  - Advantages of eBooks Over Traditional Books
2. Identifying Gsu Chemistry 1212 Lab Manual
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gsu Chemistry 1212 Lab Manual
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gsu Chemistry 1212 Lab Manual
  - Personalized Recommendations
  - Gsu Chemistry 1212 Lab Manual User Reviews and Ratings
  - Gsu Chemistry 1212 Lab Manual and Bestseller Lists
5. Accessing Gsu Chemistry 1212 Lab Manual Free and Paid eBooks
  - Gsu Chemistry 1212 Lab Manual Public Domain eBooks
  - Gsu Chemistry 1212 Lab Manual eBook Subscription Services
  - Gsu Chemistry 1212 Lab Manual Budget-Friendly Options

6. Navigating Gsu Chemistry 1212 Lab Manual eBook Formats
  - ePub, PDF, MOBI, and More
  - Gsu Chemistry 1212 Lab Manual Compatibility with Devices
  - Gsu Chemistry 1212 Lab Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gsu Chemistry 1212 Lab Manual
  - Highlighting and Note-Taking Gsu Chemistry 1212 Lab Manual
  - Interactive Elements Gsu Chemistry 1212 Lab Manual
8. Staying Engaged with Gsu Chemistry 1212 Lab Manual
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gsu Chemistry 1212 Lab Manual
9. Balancing eBooks and Physical Books Gsu Chemistry 1212 Lab Manual
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gsu Chemistry 1212 Lab Manual
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gsu Chemistry 1212 Lab Manual
  - Setting Reading Goals Gsu Chemistry 1212 Lab Manual
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gsu Chemistry 1212 Lab Manual
  - Fact-Checking eBook Content of Gsu Chemistry 1212 Lab Manual
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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