

Name _____ Date _____ Class _____

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	π bond
---------------	----------	------------	------------	------------

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5) π bond**.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
 - a single atom
 - a molecule containing three atoms**
 - a molecule containing two atoms
 - a molecule containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - four**
- Bond length is the distance between
 - two molecules of the same substance
 - the nuclei of two attached atoms**
 - the electrons in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - As the character of the bond, the stronger the bond.**
 - As the character of the bond, the weaker the bond.
 - As the character of the bond, the fewer the electrons in it.
 - As the character of the bond, the longer the bond dissociation energy.

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49

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CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write *true* or *false*.

- Binary molecular compounds are generally composed of a metal and a nonmetal.
false
- The second element in the formula of a binary compound is named using the suffix *-ite*.
false
- The prefix *tri-* indicates three atoms.
true
- The prefix *kilo-* indicates six atoms.
true
- In naming the first element in a formula, the prefix *mono-* is not used.
true
- For binary acids, the hydrogen part of the compound is named using the prefix *hydro-*.
true
- An oxyacid contains only two elements.
false
- If the name of the cation of an oxyacid ends in *-ite*, the acid name contains the suffix *-ate*.
false

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A	Column B
5. Cl_2	a. hydrobromic acid
6. CO	b. nitrogen trioxide
7. NO_2	c. carbon monoxide
8. H_2CO_3	d. nitrous acid
9. H_2SO_4	e. acetone
10. N_2O_5	f. nitric acid
11. HNO_3	g. carbonic acid
12. HNO_2	h. bromic acid
13. HBr	i. carbon dioxide
14. HBrO_3	

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Table of Contents Key Minerals Study Guide For Content Mastery

1. Understanding the eBook Key Minerals Study Guide For Content Mastery
 - The Rise of Digital Reading Key Minerals Study Guide For Content Mastery
 - Advantages of eBooks Over Traditional Books
2. Identifying Key Minerals Study Guide For Content Mastery
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Key Minerals Study Guide For Content Mastery
 - User-Friendly Interface
4. Exploring eBook Recommendations from Key Minerals Study Guide For Content Mastery
 - Personalized Recommendations
 - Key Minerals Study Guide For Content Mastery User Reviews and Ratings
 - Key Minerals Study Guide For Content Mastery and Bestseller Lists
5. Accessing Key Minerals Study Guide For Content Mastery Free and Paid eBooks
 - Key Minerals Study Guide For Content Mastery Public Domain eBooks
 - Key Minerals Study Guide For Content Mastery eBook Subscription Services
 - Key Minerals Study Guide For Content Mastery Budget-Friendly Options

6. Navigating Key Minerals Study Guide For Content Mastery eBook Formats
 - ePub, PDF, MOBI, and More
 - Key Minerals Study Guide For Content Mastery Compatibility with Devices
 - Key Minerals Study Guide For Content Mastery Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Key Minerals Study Guide For Content Mastery
 - Highlighting and Note-Taking Key Minerals Study Guide For Content Mastery
 - Interactive Elements Key Minerals Study Guide For Content Mastery
8. Staying Engaged with Key Minerals Study Guide For Content Mastery
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Key Minerals Study Guide For Content Mastery
9. Balancing eBooks and Physical Books Key Minerals Study Guide For Content Mastery
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Key Minerals Study Guide For Content Mastery
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Key Minerals Study Guide For Content Mastery
 - Setting Reading Goals Key Minerals Study Guide For Content Mastery
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Key Minerals Study Guide For Content Mastery
 - Fact-Checking eBook Content of Key Minerals Study Guide For Content Mastery
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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