

14 .2 Math Skills

$$\text{Energy} = cm\Delta t$$

$$t_f = t_i - \frac{\text{energy}}{cm}$$

$$t_f = t_i + \frac{\text{energy}}{cm}$$

$$c = \frac{\text{energy}}{m\Delta t}$$

$$m = \frac{\text{energy}}{c\Delta t}$$

$$\Delta t = \text{---}$$

$$\Delta t = t_f - t_i$$

Specific Heat

☹ To receive any credit, you must show all your work, including givens, the correct equation, givens plugged in and answer with unit.

1. Brass is an alloy made from copper and zinc. A 0.59 kg brass candlestick has an initial temperature of 98.0°C. If 2.11×10^4 J of energy is removed from the candlestick to lower its temperature to 6.8°C, what is the specific heat of brass?
2. Mercury has one of the lowest specific heats. This fact added to its liquid state at most atmospheric temperatures makes it effective for use in thermometers. If 257 J of energy are added to 0.450 kg of mercury, the mercury's temperature will increase by 4.09 K. What is the specific heat of mercury?
3. Bismuth's specific heat is 121 J/kg • K, the lowest of any non-radioactive metal. What is the mass of a bismuth sample if 25 J raises its temperature 5.0 K?

Math Skills Specific Heat

Ying-Ying Zheng

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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