

Name \_\_\_\_\_

Chemistry Period \_\_\_\_\_

### Worksheet Specific Heat

Specific Heat of water =  $4.18 \text{ J/g}^\circ\text{C}$

1. A sample of mercury is heated from  $25.5^\circ\text{C}$  to  $52.5^\circ\text{C}$ . In the process,  $3050\text{J}$  of heat are absorbed. What mass of mercury was in the sample? The specific heat of mercury is  $0.140 \text{ J/g}^\circ\text{C}$ . **ANS: 807 g**
2. A block of aluminum with a mass of  $140\text{g}$  is cooled from  $98.4^\circ\text{C}$  to  $62.2^\circ\text{C}$  with a release of  $1137\text{J}$  of heat. From these data, calculate the specific heat of aluminum. **ANS: 0.224**
3. A total of  $226 \text{ J}$  of heat are absorbed as  $58.3 \text{ g}$  of lead is heated from  $12.0^\circ\text{C}$  to  $42.0^\circ\text{C}$ . From this data, what is the specific heat of lead? **ANS: 0.129**
4. What is the change in heat energy when  $132\text{g}$  of copper at  $32.2^\circ\text{C}$  is raised to  $45.0^\circ\text{C}$ ? The specific heat of copper is  $0.385 \text{ J/g}^\circ\text{C}$ . **ANS: 650 J**
5. What is the change in heat energy when  $40.0\text{g}$  of water at  $25.0^\circ\text{C}$  is raised to  $60.0^\circ\text{C}$ ? **ANS: 5852 J**
6. What is the change in heat energy when  $75.0$  of water at  $9.8^\circ\text{C}$  is raised to  $22.4^\circ\text{C}$ ? **3950 J**
7. What is the change in heat energy when  $125\text{g}$  of water at  $10.0^\circ\text{C}$  is raised to  $90.0^\circ\text{C}$ ? **ANS: 41.8 kJ**
8. What is the change in heat energy when  $64.82 \text{ g}$  of aluminum metal at  $100.0^\circ\text{C}$  is cooled to  $82.0^\circ\text{C}$ ? The specific heat of aluminum is  $0.897 \text{ J/g}^\circ\text{C}$ . **ANS: 1047**
9. What is the change in temperature if  $24\text{g}$  of water at  $50^\circ\text{C}$  absorbs  $985 \text{ J}$  of heat energy? **ANS: 9.82**
10. What is the mass of a sample of iron if that sample lost  $2300\text{J}$  of heat energy when it cooled from  $80^\circ\text{C}$  to  $30^\circ\text{C}$ ? The specific heat of iron is  $0.449 \text{ J/g}^\circ\text{C}$ . **ANS: 102 g**
11. What is the specific heat of metal if its mass is  $26.86\text{g}$  and it requires  $418.6 \text{ J}$  of heat energy to raise its temperature from  $27.4^\circ\text{C}$  to  $67.3^\circ\text{C}$ ? **ANS: 0.391**
12. How much heat energy is needed to raise the temperature of  $59.7\text{g}$  of cadmium from  $25^\circ\text{C}$  to  $100^\circ\text{C}$ ? The specific heat of cadmium is  $0.231 \text{ J/g}^\circ\text{C}$ . **ANS: 1034 J**

# Math Skills Specific Heat Worksheet Answers

**E Durkheim**



## **Math Skills Specific Heat Worksheet Answers:**

## Whispering the Strategies of Language: An Psychological Journey through **Math Skills Specific Heat Worksheet Answers**

In a digitally-driven earth where displays reign great and quick communication drowns out the subtleties of language, the profound techniques and psychological subtleties hidden within words usually go unheard. However, located within the pages of **Math Skills Specific Heat Worksheet Answers** a charming fictional prize pulsating with natural feelings, lies an exceptional quest waiting to be undertaken. Composed by a talented wordsmith, that marvelous opus invites visitors on an introspective journey, gently unraveling the veiled truths and profound influence resonating within the fabric of every word. Within the emotional depths of this moving review, we shall embark upon a heartfelt exploration of the book is core styles, dissect their charming writing type, and yield to the strong resonance it evokes strong within the recesses of readers hearts.

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