


Circular Figures
Sphere

Cylinder

Cone

Prisms
Cube

P.S. E-12 V-8


Triangular prism

P.S. E-8 V-6


Rectangular prism

P.S. E-12 V-8


Pyramids
Square Pyramid

P.S. E-8 V-5


Triangular Pyramid

P.S. E-8 V-4


Regrouping
Addition

$$129 + 113 = 242$$

$$\begin{array}{r} 129 \\ + 113 \\ \hline 242 \end{array}$$

Subtraction

$$\begin{array}{r} 208 \\ - 33 \\ \hline 175 \end{array}$$

$$200 - 3 = 197$$

$$\begin{array}{r} 200 - 3 \\ = 200 - 1 - 2 \\ = 199 \end{array}$$

-Read every problem more than once.

-Check your answer and bubble in carefully!

STAAR Math Review
Lines
Parallel: lines that never cross

Intersecting: lines that cross

Perpendicular: lines that form right angles

Measurement
About an inch

1 foot = 12 inches

3 feet = 1 yard

Remember you can measure with a broken ruler if you count on from the start. The pencil is not 10 K inches. It's 7 K inches.

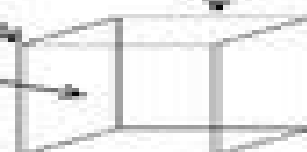

Shapes
Quadrilateral (4 sides)

Pentagon (5 sides)

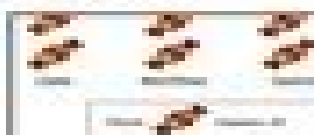
Octagon (8 sides)

Rhombus

Parallelogram

vertex
edge
face

Perimeter: the distance around an object. *the outside (add)

Area: the number of square units needed to cover a surface or shape. (*the inside)

area
Graphing
Pictograph


Check the key & labels!

Bar Graph

Check the SCALE & labels!



Math Teacher Staar Guide Practice

J Dewey



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- Interactive and Gamified eBooks

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the word titration comes from the latin word titulus which means inscription or title

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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 ml of base is added section 17 3 2 1

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0 100 m the ph of the solution is then $\text{ph} = -\log 0.100 = 1.000$ b titrant volume 12 50 ml

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web at the equivalence point in an acid base titration moles of base moles of acid and the solution only contains salt and water acid base titrations are monitored by the change of ph as titration progresses

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Volume of NaOH added (mL)	pH of the solution
4	4.6
5	4.6
6	5.9
7	5.9
8	5.9
9	5.9

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