

Lab Report: Gravimetric Analysis of an Unknown Sulfate

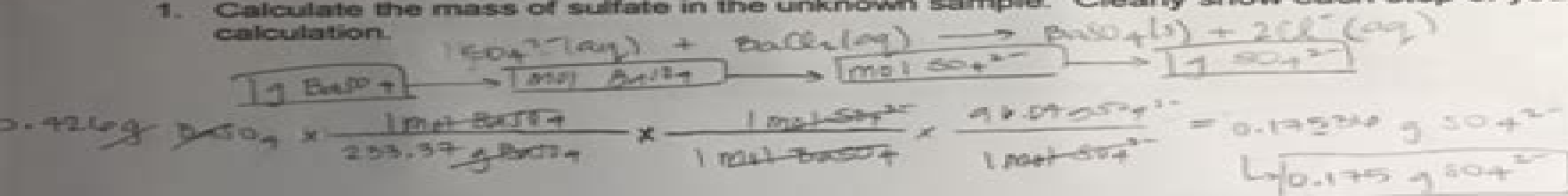
Experimental Data

Unknown Sulfate ID Code: 20

1. Mass of empty 250-mL beaker	111.244 g
2. Mass of 250-mL beaker and unknown sulfate	112.227 g
3. Mass of unknown sulfate	0.343 g
4. Mass of empty crucible (without lid)	11.203 g
5. Mass of crucible (without lid) and barium sulfate	11.629 g
6. Mass of barium sulfate	0.426 g

Calculations and Conclusions

- Calculate the mass of sulfate in the unknown sample. Clearly show each step of your calculation.



- Calculate the percentage by mass of sulfate in the unknown sample. Show your work.
mass of unknown sulfate: 0.343 g ; mass of sulfate: 0.175 g SO₄²⁻

$$\% \text{ SO}_4^{2-} = \frac{0.175 \text{ g}}{0.343 \text{ g}} \times 100 = 51.0204\%$$

$$\hookrightarrow \approx 51.0\%$$

- Calculate the percentage by mass of the metal in the unknown sample. Show your work.

$$100.0\% - 51.0\% = 49.0\%$$

Gravimetric Analysis Lab Report Ucsd

Michal Rosen-Zvi



Gravimetric Analysis Lab Report Ucsd:

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Table of Contents Gravimetric Analysis Lab Report Ucsd

1. Understanding the eBook Gravimetric Analysis Lab Report Ucsd
 - The Rise of Digital Reading Gravimetric Analysis Lab Report Ucsd
 - Advantages of eBooks Over Traditional Books
2. Identifying Gravimetric Analysis Lab Report Ucsd
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gravimetric Analysis Lab Report Ucsd
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gravimetric Analysis Lab Report Ucsd

- Personalized Recommendations
 - Gravimetric Analysis Lab Report Ucsd User Reviews and Ratings
 - Gravimetric Analysis Lab Report Ucsd and Bestseller Lists
5. Accessing Gravimetric Analysis Lab Report Ucsd Free and Paid eBooks
 - Gravimetric Analysis Lab Report Ucsd Public Domain eBooks
 - Gravimetric Analysis Lab Report Ucsd eBook Subscription Services
 - Gravimetric Analysis Lab Report Ucsd Budget-Friendly Options
 6. Navigating Gravimetric Analysis Lab Report Ucsd eBook Formats
 - ePub, PDF, MOBI, and More
 - Gravimetric Analysis Lab Report Ucsd Compatibility with Devices
 - Gravimetric Analysis Lab Report Ucsd Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gravimetric Analysis Lab Report Ucsd
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 - Interactive Elements Gravimetric Analysis Lab Report Ucsd
 8. Staying Engaged with Gravimetric Analysis Lab Report Ucsd
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gravimetric Analysis Lab Report Ucsd
 9. Balancing eBooks and Physical Books Gravimetric Analysis Lab Report Ucsd
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gravimetric Analysis Lab Report Ucsd
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gravimetric Analysis Lab Report Ucsd
 - Setting Reading Goals Gravimetric Analysis Lab Report Ucsd
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gravimetric Analysis Lab Report Ucsd

- Fact-Checking eBook Content of Gravimetric Analysis Lab Report Ucsd
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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