

**its second half equivalence point?**

- A.  $[\text{CO}_3^{2-}] = \text{maximum value}$
- B.  $[\text{HCO}_3^-] = \text{maximum value}$
- C.  $[\text{H}_2\text{CO}_3] = [\text{HCO}_3^-]$
- D.  $[\text{HCO}_3^-] = [\text{CO}_3^{2-}]$
- E.  $[\text{H}_2\text{CO}_3] = [\text{CO}_3^{2-}]$

Answer: D.  $[\text{HCO}_3^-] = [\text{CO}_3^{2-}]$

**4. The  $K_a$  for HCN is  $6.2 \times 10^{-10}$ , If there is a solution of 2M HCN, what concentration of NaCN is needed in order for the pH to be 9.2?**

- A. 3M
- B. No NaCN needs to be added
- C. 2M
- D. 1M

Answer: C. 2M

**5. NaOH is added to a 500ml of 2M acetic acid. If the  $\text{p}K_a$  value of acetic acid is approximately 4.8, what volume of 2M NaOH must be added so that the pH of the solution is 4.8?**

- A. 2L
- B. 500ml
- C. 250mL
- D. 1L

Answer: C. 250ml

**6. Consider the titration curve, in which an acid is treated with aqueous 0.5M NaOH. What is the approximate  $\text{p}K_a$  of the starting acid?**

- A. 4.2
- B. 3.0
- C. 8.9
- D. 5.0
- E. 7.0

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