

**Assessment 1****Chapter Test B****Chapter: Chemical Equations and Reactions**

**Part 1** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The production of a slightly soluble solid compound in a double-displacement reaction is called the formation of a  
a. gas.  
b. precipitate.  
c. molecular reaction.  
d. halogen.
- \_\_\_\_\_ 2. To balance a chemical equation, it is permissible to adjust the  
a. coefficients.  
b. subscripts.  
c. formulae of the products.  
d. number of products.
- \_\_\_\_\_ 3. In a chemical equation, the symbol (aq) indicates that the substance is  
a. melted.  
b. dissolved in water.  
c. an acid.  
d. insoluble.
- \_\_\_\_\_ 4. The tendency for a replacement reaction to occur depends on the  
a. relative reactivities of the elements in the activity series of metals.  
b. temperature difference.  
c. valence electrons and mass.  
d. relative reactivities of the elements in the activity series of metals.
- \_\_\_\_\_ 5. The coefficients in a chemical equation  
a. indicate the number of moles of each substance that react.  
b. show the number of grams of each substance that react.  
c. are the molar masses of the substances.  
d. show the valence electrons for each atom.
- \_\_\_\_\_ 6. If metal X is more reactive than metal Y in the activity series, then  
a. X will replace ions of Y in solution.  
b. Y will replace ions of X in solution.  
c. Y will react with cold water.  
d. X will react with cold water.

# Guide Key Chemistry

**Siddappa N.Byrareddy**



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