AP Biology Study Guide Unit 1

Properties of Water:

- 1. Polarity allows H20 to form Hydrogen Bonds
- 2. Cohesion Water H bonds to itself, creating surface tension
- 3. Adhesion Water adheres to other polar substances
 - a. Allows easy transport of water against Gravity
- 4. Universal Solvent Water can dissolve Polar and Ionic Compounds
 - a. This allows for reactions between Aqueous Solutions
- 5. Expands when Frozen this means that Ice floats
- High Specific Heat water can absorb or release a large amount of heat with only a slight change in its own temperature (caused by cohesion)
 - High Heat of Vaporization only warmest molecules can evaporate, so this causes evaporative cooling
- 7. Can be regulated by non-polar hydrophobic substances
- Neutral PH → can form acids or bases if substances are added

Polarity - Difference in electronegativity between atoms bonded together create slightly charged ends of the molecules they form

Hydrophobic Substance - Nonpolar substance that repels water Hydrophilic Substance - Polar substance that attracts water molecules

CHON - Carbon, Hydrogen, Nitrogen, Oxygen

- Makeup 99% of living matter
- Strong Biological Cornerstones Because:
 - Smallest Diameter/Valency for atoms of 1, 2, 3, 4 valence electrons
 - Form Stable Covalent Bonds
 - Can form double and triple bonds
 - Versatile

Biological Transition Metals - Mn. Fe. Co. Cu. Zn.

Can exist in multiple oxidation states → can be both oxidized and reduced

Biological Ions: Na+, Mg2+, Cl-, K+, Ka2+

Used to maintain homeostasis and osmotic concentration

Phosphorous/Sulphur - Energy Carriers which form strong high energy bonds

5 Foundation Biological Molecules:

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Lauren Gardner

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Table of Contents Guide Line Study For Ap Biology Campbell

- 1. Understanding the eBook Guide Line Study For Ap Biology Campbell
 - The Rise of Digital Reading Guide Line Study For Ap Biology Campbell
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Guide Line Study For Ap Biology Campbell
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guide Line Study For Ap Biology Campbell
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Guide Line Study For Ap Biology Campbell
 - Personalized Recommendations
 - Guide Line Study For Ap Biology Campbell User Reviews and Ratings
 - Guide Line Study For Ap Biology Campbell and Bestseller Lists
- 5. Accessing Guide Line Study For Ap Biology Campbell Free and Paid eBooks
 - Guide Line Study For Ap Biology Campbell Public Domain eBooks
 - Guide Line Study For Ap Biology Campbell eBook Subscription Services

- Guide Line Study For Ap Biology Campbell Budget-Friendly Options
- 6. Navigating Guide Line Study For Ap Biology Campbell eBook Formats
 - ∘ ePub, PDF, MOBI, and More
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 - Guide Line Study For Ap Biology Campbell Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guide Line Study For Ap Biology Campbell
 - Highlighting and Note-Taking Guide Line Study For Ap Biology Campbell
 - Interactive Elements Guide Line Study For Ap Biology Campbell
- 8. Staying Engaged with Guide Line Study For Ap Biology Campbell
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guide Line Study For Ap Biology Campbell
- 9. Balancing eBooks and Physical Books Guide Line Study For Ap Biology Campbell
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guide Line Study For Ap Biology Campbell
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Guide Line Study For Ap Biology Campbell
 - Setting Reading Goals Guide Line Study For Ap Biology Campbell
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guide Line Study For Ap Biology Campbell
 - Fact-Checking eBook Content of Guide Line Study For Ap Biology Campbell
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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