# Chemistry Lab



#### Introduction to working as a Chemist.

Use a highlighter as you work through this booklet. ESL students, make vocab lists of the terminology that is defined or described here.

#### Assessment Statements

Outline the Scientific Method State SI units of measurement for Identify independent, dependent & controlled Mass: grams (q) variables. Length: Metres (m), millimeters (mm) Demonstrate correct and safe use of laboratory Concentration: molarity (M) equipment. Temperature: Degrees Celsius (°C) Distinguish between quantitative and qualitative Volume: Litres (L) or millitres (mL) Density: grams per cubic centimeter Qualitative is descriptive, not using Eq./cmi\*) numbers, e.g. properties, colour, gas Conductivity: Siemens (S) production, sound, odor. Acidity or alkalinity (pH) Quantitative is mathematical, including Energy: Joules (J or kJ) Determine the Raw data (as recorded) uncertainty of digital and analog measuring tools Processed (after calculations) Usually ± 1 of the smallest division Distinguish between accuracy and precision Calculate the % error of a given or recorded set Accuracy describes the correctness of a cold words were: set of measurements ((Measured - True) / True) x 100 Precision describes the repeatability of a State values to appropriate numbers of set of measurements significant figures Calculate values to appropriate numbers of significant figures

#### Related Concepts:

□ 01.1 Element symbols (Quia quie: http://www.guia.com/quie/3246640.html)
□ 01.2 Accuracy, Precision, Error, Uncertainty (Quia quie: http://www.guia.com/quie/3507626.html)
□ 01.3 SigDigs and Notation (Quia quie: http://www.guia.com/quie/3504378.html)

#### Command Terms: Be sure you understand what is being asked of you

Analyse Break down to identify key parts. Interpret information to reach conclusions. Apply Use knowledge & understanding in response to a real situation or given circumstances. Calculate: Obtain a numerical answer, showing relevant stages in the working. Define. Give the precise meaning of a word, phrase, concept or physical quantity. Give a detailed account or picture of an event, pattern, process or pattern. Describe Design Produce a plan, simulation or model. Dieter milme Obtain the only possible answer. Discuss Offer a considered and balanced review that includes a range of arguments, factors or hypotheses. Opinions or conclusions should be presented clearly and supported by appropriate evidence. Distinguish Make clear the differences between two or more concepts or items. Evaluate: Assess the implications and limitations: make judgments about the ideas, works, solutions or ideas in relation to selected criteria. Explain Give a detailed account including reasons or causes. Identify. Provide an answers from a number of possibilities. Recognise and state briefly a distinguishing fact or feature. Add title, labels or brief explanation(s) to a diagram or graph. Lobel List Give a sequence of answers with no explanation. Give a brief account. Outline Recall Remember or recognize from prior learning experiences. Store Give a specific name, value or other brief answer without explanation or calculation. Suggest Propose a solution, hypothesis or other possible answer.

# **Lab 111 Laboratory Manual Answers**

Mary Alice Minor, Christopher Towler, Lynn S. Lippert

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stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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