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LAB MANUAL

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MECHANICAL ENGINEERING DEPARTMENT

**KHWAJA FAREED UNIVERSITY OF ENGINEERING &
TECHNOLOGY RAHIM YAR KHAN**

Lab Manual For Civil Diploma Engg Cad

Tobias Bleicker



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subsequent Δh values $2 \text{ CO}_2 \text{ g} + \text{H}_2\text{O} \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{O}_2 \text{ g}$ $\Delta h = 2 \text{ H}_2 \text{ g} + \text{C}_2\text{H}_6 \text{ g} \rightarrow \Delta h = 94.5 \text{ kJ}$ $\text{H}_2\text{O} \text{ g} \rightarrow \text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g}$ $\Delta h = 71.2 \text{ kJ}$ $\text{C}_2\text{H}_6 \text{ g} \rightarrow \frac{7}{2} \text{O}_2 \text{ g} + 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O} \text{ g}$ $\Delta h = 283 \text{ kJ}$ g $1 \text{ C}_2\text{H}_6 \text{ g}$

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changes of its individual steps hess s law example 12.1 determine Δh for the target reaction $\text{C}_2\text{H}_4 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2\text{O} \text{ g}$
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CO_2 g 2 H_2O l h 1411 kJ C_2H_6 g $3\frac{1}{2}$ O_2 g 2 CO_2 g 3 H_2O l h 1560 kJ

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