

Experiment 14

Data and Calculations: Heat Effects and Calorimetry

A. Specific Heat

| | Trial 1 | | Trial 2 |
|---|---------------------------------|---|-----------------|
| Mass of stoppered test tube plus metal | <u>45.393</u> g | → | <u>45.406</u> g |
| Mass of test tube and stopper | <u>22.525</u> g | → | <u>22.539</u> g |
| Mass of calorimeter | <u>3.978</u> <u>47.173</u> g | → | <u>8.608</u> g |
| Mass of calorimeter and water | <u>45.798</u> g | | <u>48.295</u> g |
| Mass of water | <u>39.175</u> g | | <u>39.687</u> g |
| Mass of metal | <u>22.868</u> g | → | <u>22.868</u> g |
| Initial temperature of water in calorimeter | <u>21.2</u> °C | | <u>21.9</u> °C |
| Initial temperature of metal (assume 100°C unless directed to do otherwise) | <u>99.8</u> °C | → | <u>100.3</u> °C |
| Equilibrium temperature of metal and water in calorimeter | <u>25.3</u> °C | | <u>24.5</u> °C |
| $\Delta t_{\text{water}} (t_{\text{final}} - t_{\text{initial}})$ | _____ °C | | _____ °C |
| Δt_{metal} | _____ °C | | _____ °C |
| $q_{\text{H}_2\text{O}}$ | _____ J | | _____ J |
| Specific heat of the metal (Eq. 3) | _____ J/g°C | | _____ J/g°C |
| Approximate molar mass of metal | _____ | | _____ |
| Unknown metal # | | | <u>26Y</u> |

B. Heat of Solution

| | |
|--------------------------------|-----------------|
| Mass of calorimeter plus water | <u>56.898</u> g |
| Mass of beaker | <u>29.309</u> g |

(continued on following page)

Lab Manual Chemistry Heat Effects And Calorimetry

Marcel A. Müller



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