

Procedure

You will use the simulation on the following website to virtually measure and record all voltages for this experiment.

<http://web.mst.edu/~gbert/Electro/Electrochem.html>

Once you have gone to the website above you should see the interface below.



- You will be able to adjust the electrodes and solutions using the drop down menus on the right and left. The simulator will only allow you to use a solution of the metal salt corresponding to the metal electrode placed in the beaker. (For example: Copper(II) nitrate must be used with the copper electrode.)
- You can also adjust the concentration of each of the metal salt solutions in the beakers by typing the value in the box on either side of the simulator.
- To measure the voltage of any cell that you have constructed, simply click **Measure Cell Voltage** on the bottom left of the simulator. If there is a problem with the cell, the simulator will identify any errors in your construction.
- You will not be using the **New Problem** link in the or "Level" drop down menu in the lower right for this assignment.

Part A

1. You will be measuring the cell potentials for all the possible combinations of each of the following electrodes: Ag/Ag^+ , Cu/Cu^{2+} , Ni/Ni^{2+} , Pb/Pb^{2+} , and Zn/Zn^{2+} . Construct the cells so that the metal salt solution in each beaker has a concentration of 0.10 M and a positive voltage is read on the voltmeter when the cell potential is measured.
2. Place a silver electrode in the beaker on the left using the top drop down menu and then select silver nitrate in the second drop down menu and type in a concentration of 0.10 into the box at the bottom.

Lab Manual For Chem 1212l

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A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the American Chemical Society text.

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CHEM 1212 LAB, KSU, Huck Benjamin Huck, N. J. Kirk, 2014-01-01 Chemistry Lab Manual **Chemistry 1212 L Laboratory Manual**, 2015-12 *Laboratory Experiments in General Chemistry* Alan M. Gabrielli, Jack Duff, 2006
Laboratory Manual for Chemical and Bacterial Analysis of Water and Sewage Frank Richard Theroux, Edward F. Eldridge, Walter LeRoy Mallmann, 1936 **A Laboratory Manual for the Chemical Analysis of Water and Sewage** Edward F. Eldridge, Frank Richard Theroux, 1935 *Course and Curriculum Improvement Projects: Mathematics, Science, Social Sciences* National Science Foundation (U.S.), 1970 **Journal of the American Medical Association**, 1928
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- Chapter 5: Conclusion

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