

## Chem lab report

### Design

**Aim:** To determine the moles of water of crystallization in oxalic acid.

**Hypothesis:** If NaOH is added to Hydrated Oxalic acid, then the hydrogen ions from the acid will react with hydroxide ions from NaOH to produce water. The rest forms a colorless salt.

In this experiment, the moles of water of crystallization is calculated using titration. An indicator is used to show when the titration has ended. Since two moles of NaOH react with one mole of oxalic acid, we can find out the moles of oxalic acid used, then using the mass of the crystals taken, find the mass of water present in the crystals which will give the molecules of water of crystallization in oxalic acid.

### Variables:

- Independent :
  - Volume of NaOH used- since  $25\text{cm}^3$  of NaOH is taken for all trials
- Dependent:
  - Volume of oxalic acid solution used-the volume of oxalic acid solution used would

# Lab 57 Titration Oxalic Acid

**J Spring**



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*Chemistry* Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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National Laboratory, 1959      *Chemistry Lab Manual* Neena Sinha, R Rangarajan, R P Manchanda, R K Gupta, Rajesh Kumar,

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