

Assessment 3**Chapter Test B****Chapter: Chemical Equations and Reactions**

Part 1 In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The production of a slightly soluble solid compound in a double-displacement reaction is called the formation of a
a. gas.
b. precipitate.
c. combustion reaction.
d. halogen.
- _____ 2. To balance a chemical equation, it is permissible to adjust the
a. coefficients.
b. subscripts.
c. formulae of the products.
d. number of products.
- _____ 3. In a chemical equation, the symbol (aq) indicates that the substance is
a. melted.
b. dissolved in water.
c. an acid.
d. insoluble.
- _____ 4. The tendency for a replacement reaction to occur depends on the
a. difference between any two elements in the activity series themselves.
b. temperature difference.
c. valence electrons and mass of.
d. difference between any two elements in the activity series themselves.
- _____ 5. The coefficients in a chemical equation
a. indicate the number of moles of each substance that react.
b. show the number of grams of each substance that react.
c. are the molar masses of the substances.
d. show the valence electrons for each atom.
- _____ 6. If metal X is more reactive than metal Y in the activity series, then
a. X will replace ions of Y in solution.
b. Y will replace ions of X in solution.
c. Y will react oxidize only indirectly.
d. X will react with cold water.

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Michael Hoelscher



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R R Bowker Publishing,Bowker,1999-12 **Children's Books in Print, 2007** ,2006 Forthcoming Books Rose
Army,2003-04 Chapter Resource 2 Chemistry of Life Biology Holt Rinehart & Winston,Holt, Rinehart and Winston
Staff,2004 **Economics and Thermodynamics** Peter Burley,John Foster,2013-03-09 Over the past two decades we have
witnessed something of a revolution in the natural sciences as thermodynamic thinking evolved from an equilibrium or
classical perspective to a nonequilibrium or self organisational one In this transition thermodynamics has been applied in
new ways and in new fields of inquiry Chemical and biological evolutionary processes have been analysed increasingly in non
equilibrium thermodynamical terms Economics has since the late 19th century relied heavily upon metaphors and analogies
derived from the natural sciences mechanical analogies cast in terms of traditional Newtonian physics and expressed in
terms of Cartesian logic have been especially popular Thermodynamics on the other hand has been less popular despite its
early application in economics by Stanley Jevons the father of modern notions of utility maximisation in neoclassical
economics and despite its promotion in economic contexts by Paul Samuelson the author of the definitive treatise upon which
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thermodynamic thinking in economics was brought to our attention by Nicholas Georgescu Roegen in the late 1960s by
which time economic theory evidenced in for example the Arrow Debreu general equilibrium system had become so
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crisis in economics because neglect of thermodynamics led in his view to blindness amongst economists to an economy
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