

Teacher Guide: Calorimetry Lab



Learning Objectives

Students will ...

- Use a calorimeter to examine how heat energy transfers from a material to water.
- Compare the specific heat capacities of various materials.
- Determine the amount of energy released by cooling materials.
- Calculate the specific heat capacities of various materials.



Vocabulary

calorie, calorimeter, joule, specific heat capacity



Lesson Overview

Materials differ in how easily they change temperature when they absorb or lose energy. This property is known as *specific heat capacity*. The specific heat capacity of a substance is the amount of energy needed to increase that substance's temperature by 1 °C.

Calorimeters are devices that can be used to determine an object's specific heat capacity. Students can use the *Calorimetry Lab Gizmo* to calculate and compare the specific heat capacities of several different materials.



The Student Exploration sheet contains three activities:

- **Activity A** – Students investigate the factors that affect how heat transfers between objects.
- **Activity B** – Students compare the specific heat capacities of various substances.
- **Activity C** – Students use the specific heat equation to determine how much heat energy a substance gains or loses and to calculate the specific heat capacity of a substance.



Suggested Lesson Sequence

1. **Pre-Gizmo activity: Measure temperature change** (🕒 10 – 15 minutes)
Pour 250 g (250 mL) of pure water into a beaker. Pour 250 g (about 272 mL) of vegetable oil into a second beaker. Place both beakers over Bunsen burners or hot plates. Have a student assistant stir each substance as you take their temperatures every 10–20 seconds. Have another student record each substance's temperature readings on the board. After about 5 minutes, have students compare the temperature changes of the two substances.

After doing the experiment, ask students to identify the substance that heated up the fastest. Assuming the Bunsen burners are at the same intensity, the vegetable oil should heat up about twice as quickly as the water. Have students discuss why they think this may have occurred.

Instructors Guide Calorimetry

Joacim Rocklöv



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