



# CHEMISTRY

## CHEAT SHEET

# EQUILIBRIUM

## EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol  $\rightleftharpoons$ .

<p><b>What is Equilibrium?</b></p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In chemistry, it refers to a reaction where the forward and reverse rates are equal, and the concentrations of reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the net change is zero.</p>	<p><b>Equilibrium Constant</b></p> <p>The equilibrium constant (<math>K</math>) is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	
<p><b>Equilibrium Constant, <math>K_c</math></b></p> <p><math>K_c</math> is the equilibrium constant based on concentrations. It is used for reactions in solution or in a closed system where the concentrations of the species are known.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant <math>K_c</math> is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Equilibrium Constant, <math>K_p</math></b></p> <p><math>K_p</math> is the equilibrium constant based on partial pressures. It is used for reactions involving gases, where the partial pressures of the species are known.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant <math>K_p</math> is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	
<p><b>Equilibrium Constant, <math>K</math></b></p> <p><math>K</math> is the general equilibrium constant, which can be either <math>K_c</math> or <math>K_p</math>. It is used to describe the equilibrium of a reaction in any system.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant <math>K</math> is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Relationship between <math>K_c</math> and <math>K_p</math></b></p> <p>The relationship between <math>K_c</math> and <math>K_p</math> is given by the equation:</p> $K_p = K_c (RT)^{\Delta n}$ <p>where <math>R</math> is the gas constant, <math>T</math> is the temperature in Kelvin, and <math>\Delta n</math> is the change in the number of moles of gas.</p>	
<p><b>Le Chatelier's Principle: the effect of changing conditions on equilibrium</b></p>		
<p><b>Concentration</b></p>	<p>• Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products.</p> <p>• Increasing the concentration of a product will shift the equilibrium to the left, favoring the reactants.</p>	<p><math>N_2 + 3H_2 \rightleftharpoons 2NH_3</math></p> <p>Increasing the concentration of <math>N_2</math> will shift the equilibrium to the right, producing more <math>NH_3</math>.</p>
<p><b>Pressure</b></p>	<p>• Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas.</p> <p>• Decreasing the pressure will shift the equilibrium to the side with the most moles of gas.</p>	<p><math>N_2 + 3H_2 \rightleftharpoons 2NH_3</math></p> <p>Increasing the pressure will shift the equilibrium to the right, favoring the products (<math>2NH_3</math>).</p>
<p><b>Temperature</b></p>	<p>• Increasing the temperature will shift the equilibrium to the side that is endothermic (absorbs heat).</p> <p>• Decreasing the temperature will shift the equilibrium to the side that is exothermic (releases heat).</p>	<p><math>N_2 + 3H_2 \rightleftharpoons 2NH_3</math></p> <p>Increasing the temperature will shift the equilibrium to the left, favoring the reactants (<math>N_2</math> and <math>H_2</math>).</p>

# Guide Chemical Equilibrium

**Tao Wei**



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