

EXPERIMENT NO.1 **EXPERIMENT NAME: Redox Titration**

Objective: - Find out the percentage of iron in a given iron alloy solution, when X grams of alloy dissolved in 1 liter

Apparatus/Reagent Required :- Burette, Pipette, Conical flask, Funnel,
Std KMnO_4 solution, KMnO_4 solution, dil H_2SO_4 , Alloy solution.

Theory: -

To determine the amount of iron in alloy, the alloy sample is dissolved in dil. H_2SO_4 , the iron present is converted to ferrous sulfate with the liberation of hydrogen gas.



The Fe^{2+} present in the solution can be determined by titration with a standard KMnO_4 solution. Potassium permanganate oxidizes ferrous ion into ferric ion in acid medium in cold and is itself reduced to colourless manganous ion.



The appearance of a faint but permanent pink colour marks the end point.

Chemical Equation:-



Procedure: -

- Fill burette with KMnO_4 solution.

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