

Chemistry 100: Final Exam Questions

Question 1: Which compound is a substituted alkane?
 a. ethane, CH_3CH_3
 b. propene, $\text{CH}_3\text{CH}=\text{CH}_2$
 c. butanol, $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$
 d. acetone, CH_3COCH_3
 e. benzene, C_6H_6

Question 2: Which of the following is the correct Lewis structure for water?
 a. $\text{H}-\text{O}-\text{H}$
 b. $\text{H}-\text{O}-\text{H}$ with two lone pairs on oxygen
 c. $\text{H}-\text{O}-\text{H}$ with one lone pair on oxygen
 d. $\text{H}-\text{O}-\text{H}$ with three lone pairs on oxygen

Question 3: Which of the following represents the correct Lewis structure for carbon dioxide?
 a. $\text{O}=\text{C}=\text{O}$
 b. $\text{O}=\text{C}=\text{O}$ with four lone pairs on each oxygen
 c. $\text{O}=\text{C}=\text{O}$ with two lone pairs on each oxygen
 d. $\text{O}=\text{C}=\text{O}$ with one lone pair on each oxygen

Question 4: A gas sample weighs 0.450 g. It has a volume of 100 mL at 25°C and 1 atm. The gas is composed of 80% carbon and 20% hydrogen. What is the molecular formula of the gas?
 a. C_2H_4
 b. C_2H_6
 c. C_3H_8
 d. C_4H_{10}
 e. C_5H_{12}

Question 5: The compound with the empirical formula CH_2O has a molar mass of 60 g/mol. What is the molecular formula?
 a. CH_2O
 b. $\text{C}_2\text{H}_4\text{O}_2$
 c. $\text{C}_3\text{H}_6\text{O}_3$
 d. $\text{C}_4\text{H}_8\text{O}_4$
 e. $\text{C}_5\text{H}_{10}\text{O}_5$

Question 6: A sample of pure carbon dioxide gas is collected at 25°C and 1 atm. The gas has a mass of 1.00 g. What is the volume of the gas in liters?
 a. 0.50 L
 b. 1.00 L
 c. 1.50 L
 d. 2.00 L
 e. 2.50 L

Question 7: Which of the following represents the correct Lewis structure for ammonia?
 a. $\text{H}-\text{N}-\text{H}$
 b. $\text{H}-\text{N}-\text{H}$ with one lone pair on nitrogen
 c. $\text{H}-\text{N}-\text{H}$ with two lone pairs on nitrogen
 d. $\text{H}-\text{N}-\text{H}$ with three lone pairs on nitrogen
 e. $\text{H}-\text{N}-\text{H}$ with four lone pairs on nitrogen

Question 8: The density of a gas is 1.25 g/L at 25°C and 1 atm. The gas is composed of 80% carbon and 20% hydrogen. What is the molecular formula of the gas?
 a. C_2H_4
 b. C_2H_6
 c. C_3H_8
 d. C_4H_{10}
 e. C_5H_{12}

Question 9: The density of a gas is 1.50 g/L at 25°C and 1 atm. The gas is composed of 80% carbon and 20% hydrogen. What is the molecular formula of the gas?
 a. C_2H_4
 b. C_2H_6
 c. C_3H_8
 d. C_4H_{10}
 e. C_5H_{12}

Question 10: A gas sample weighs 0.500 g. It has a volume of 100 mL at 25°C and 1 atm. The gas is composed of 80% carbon and 20% hydrogen. What is the molecular formula of the gas?
 a. C_2H_4
 b. C_2H_6
 c. C_3H_8
 d. C_4H_{10}
 e. C_5H_{12}

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