



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Chemical reactions are represented by the reaction: $aA + bB \rightleftharpoons cC + dD$

What is Equilibrium? Equilibrium is a state of balance between two or more opposing processes. In chemistry, it refers to a state where the concentrations of the reactants and products remain constant over time. At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction. Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the concentrations of the reactants and products are constant.	Equilibrium Constant The equilibrium constant, K, is a measure of the extent to which a chemical reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For the reaction: $aA + bB \rightleftharpoons cC + dD$ $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products, respectively.	Graph showing concentration vs. time for a reaction at equilibrium. Reactant concentration: $[A]_0 - [A]_t = [A]_{eq}$ Product concentration: $[C]_0 + [C]_t = [C]_{eq}$
Equilibrium Constant, K_c K_c is the equilibrium constant expressed in terms of molar concentrations. It is used for reactions in which the reactants and products are in the same phase (usually aqueous or gaseous). For the reaction: $aA + bB \rightleftharpoons cC + dD$ $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where $[A]$, $[B]$, $[C]$, and $[D]$ are the molar concentrations of the reactants and products, respectively.	Equilibrium Constant, K_p K_p is the equilibrium constant expressed in terms of partial pressures. It is used for reactions in which the reactants and products are gases. For the reaction: $aA + bB \rightleftharpoons cC + dD$ $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$ where P_A, P_B, P_C, and P_D are the partial pressures of the reactants and products, respectively.	
Equilibrium Constant, K K is the general equilibrium constant, which can be expressed in terms of either concentrations (K_c) or partial pressures (K_p). For the reaction: $aA + bB \rightleftharpoons cC + dD$ $K = \frac{[C]^c [D]^d}{[A]^a [B]^b} \quad \text{or} \quad K = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	Relationship between K_c and K_p The relationship between K_c and K_p is given by the equation: $K_p = K_c (RT)^{\Delta n}$ where R is the gas constant, T is the temperature in Kelvin, and Δn is the change in the number of moles of gas.	
Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Condition	Effect on Equilibrium	Result
Concentration	- Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products. - Increasing the concentration of a product will shift the equilibrium to the left, favoring the reactants.	For the reaction: $aA + bB \rightleftharpoons cC + dD$ Increasing $[A]$ or $[B]$ will increase $[C]$ and $[D]$. Increasing $[C]$ or $[D]$ will decrease $[A]$ and $[B]$.
Pressure	- Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas. - Decreasing the pressure will shift the equilibrium to the side with the most moles of gas.	For the reaction: $aA + bB \rightleftharpoons cC + dD$ If $a + b > c + d$, increasing pressure will favor the products. If $a + b < c + d$, increasing pressure will favor the reactants.
Temperature	- Increasing the temperature will shift the equilibrium to the side that is endothermic (absorbs heat). - Decreasing the temperature will shift the equilibrium to the side that is exothermic (releases heat).	For the reaction: $aA + bB \rightleftharpoons cC + dD$ If the reaction is endothermic ($\Delta H > 0$), increasing temperature will favor the products. If the reaction is exothermic ($\Delta H < 0$), increasing temperature will favor the reactants.

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