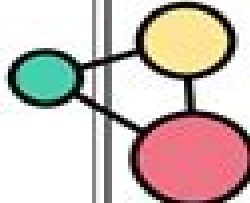
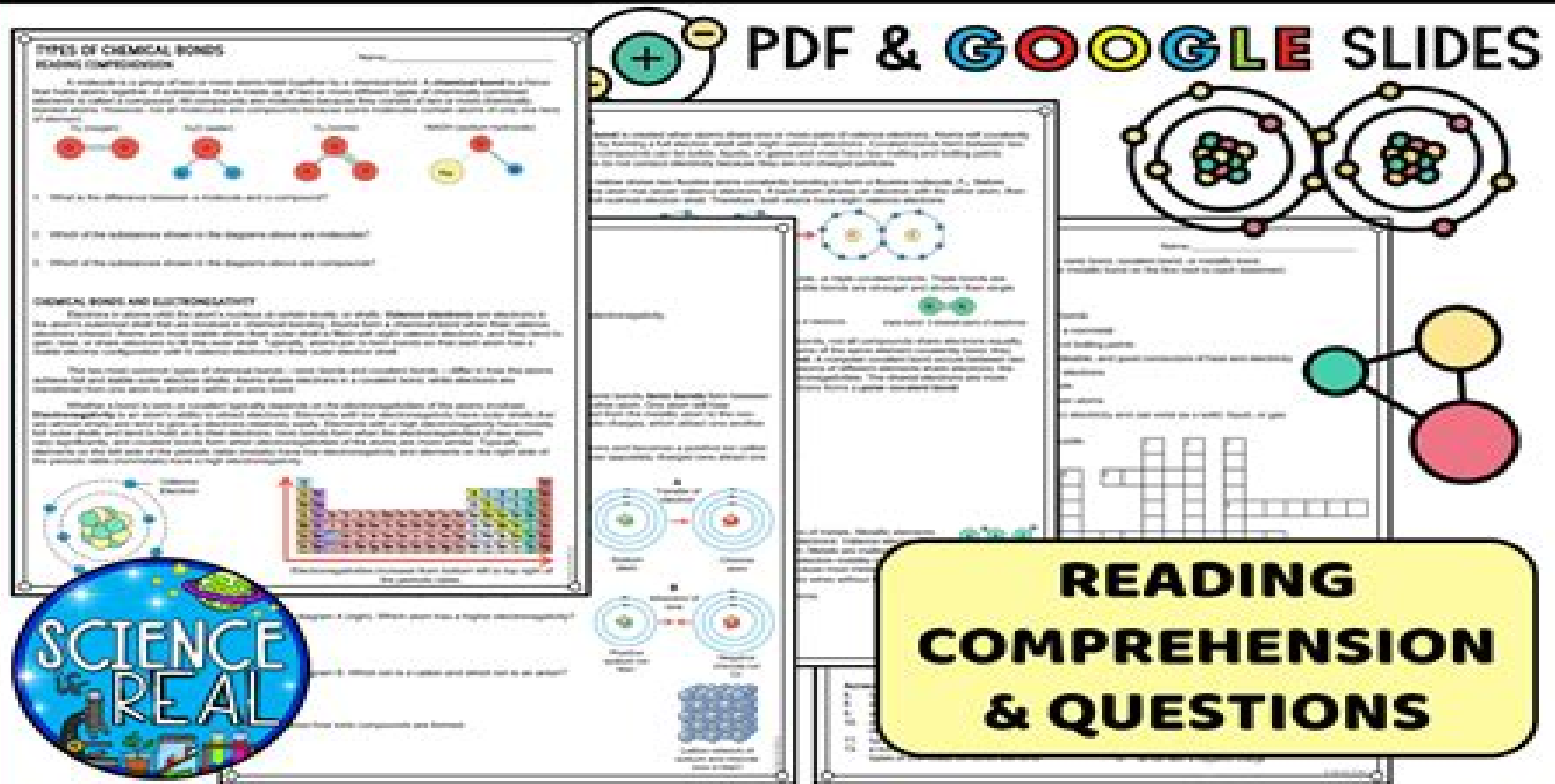


CHEMICAL BONDS

IONIC, COVALENT, & METALLIC

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TYPES OF CHEMICAL BONDS
READING COMPREHENSION

A molecule is a group of two or more atoms held together by a chemical bond. A chemical bond is a force that holds atoms together. It is sometimes thought of as a glue that holds atoms together. There are three main types of chemical bonds: ionic, covalent, and metallic. Each type of bond has its own characteristics and is found in different substances.

1. What is the difference between a molecule and a compound?

2. Which of the substances shown in the diagrams above are molecules?

3. Which of the substances shown in the diagrams above are compounds?

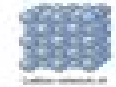

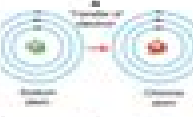
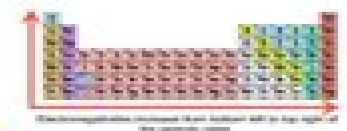

CHEMICAL BONDS AND ELECTRONEGATIVITY

Electronegativity is a measure of an atom's tendency to attract electrons. In a chemical bond, the more electronegative atom will attract the electrons more strongly than the less electronegative atom. This difference in electronegativity is what determines the type of bond that will form between two atoms.


The two most common types of chemical bonds are ionic bonds and covalent bonds. Ionic bonds form between a metal and a nonmetal, while covalent bonds form between two nonmetals.

Whether a bond is ionic or covalent typically depends on the electronegativity of the atoms involved. Electronegativity is a property of an atom that is related to its ability to attract electrons. Atoms with a high electronegativity have a strong tendency to attract electrons, while atoms with a low electronegativity have a weak tendency to attract electrons.

Electronegativity increases from left to right across the periodic table. This is because the number of protons in the nucleus increases, which increases the positive charge of the nucleus. This positive charge attracts the negatively charged electrons more strongly, which increases the electronegativity of the atom.



READING COMPREHENSION & QUESTIONS



GUIDED READING ACTIVITY

Guided Answers Ionic Bonds Ionic Compounds

**Maryanne Baudo, N.P-C ; M.S.N;
R.N.,Robin Kavanagh**



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Student's Guide to Fundamentals of Chemistry Jo A. Beran,2013-10-22 Student s Guide to Fundamentals of Chemistry Fourth Edition provides an introduction to the basic chemical principles This book deals with various approaches to chemical principles and problem solving in chemistry Organized into 25 chapters this edition begins with an overview of how to define and recognize the more common names and symbols in chemistry This text then discusses the historical development of the concept of atom as well as the historical determination of atomic weights for

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