



MONOHYBRID CROSS WORKSHEET 3

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Name _____

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Monohybrid cross problems

1. A gene encoding mouse coat color has two different alleles that either encode gray (G) or white (g) fur, or dominant and recessive traits, respectively.

a. What would the genotypes and phenotypes of the F1 generation be if a heterozygous gray mouse is mated with a white mouse?

P1 cross: Gg x gg

F1 genotype(s): 1/2 Gg, 1/2 gg

F1 phenotypes: 1/2 gray; 1/2 white

	G	g
g	Gg	gg
g	Gg	gg

b. What would the F1 genotypes and phenotypes be if a homozygous gray mouse mates with a white mouse?

P1 cross: GG x gg

F1 genotype(s): All Gg

F1 phenotypes: All gray

	G	G
g	Gg	Gg
g	Gg	Gg

2a. Two mice with unknown coat colors mated and produced the following offspring: 8 gray coated- and 4 white-coated mice. What were the genotypes and phenotypes of the parents?

P1: Gg x Gg, both gray.

[Gg x Gg] produces 1/4 gg, 1/2 Gg, 1/4 GG. Out of 12 siblings, ((1/4 + 1/2) = 3/4), or 8, are expected to be GG or Gg, or gray. 1/4 of 12, or 4, are expected to be gg, or white.

2b. If the parents produced all gray offspring, what were the genotypes and phenotypes of the parents?

P1: GG x gg, gray and white.

2c. If 50% of the offspring were gray, what were the phenotypes and genotypes of the parents?

P1: Gg x gg, gray and white.

3. A species of orchids have a gene encoding either a dominant pink (P) or recessive white (p) flower color trait. If a heterozygous pink and white flower were crossed, what is the probability that the offspring have white flowers?

50%



Life Science Genetics Monohybrid Crosses Worksheet

Rachel S Tattersall



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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
8.3 mole H₂S \times 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N
what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$ if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ grams O}_2}{32 \text{ grams/mole}} = 0.6656 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design
yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
4 Al + 3 O₂ \rightarrow 2 Al₂O₃
b determine the theoretical yield of Al₂O₃
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.926 \text{ mol Al}$
 $\frac{5.926 \text{ mol Al}}{4} \times 2 = 2.963 \text{ mol Al}_2\text{O}_3$
 $2.963 \text{ mol Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 302.3 \text{ g Al}_2\text{O}_3$
c determine the percent yield
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ label 3 7 3

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