

### Unit 3 Review - Atomic Structure & Periodic Table

#### Part 1: Answer the following questions.

1. a. Which scientist created the first modern atomic theory? John Dalton  
 b. What was his theory?
  1. All matter is made of extremely small particles called atoms
  2. All atoms of a given element are identical and atoms of a specific element are different from any other element
  3. Atoms cannot be created, divided into smaller particles or destroyed
  4. Different atoms combine in simple whole number ratios to form compounds
  5. In a chemical reaction, atoms are separated, combined or rearranged.
- c. What was his model of the atom called? Billiard Ball
- d. What was proven wrong and why?
  1. All atoms of a given element are identical – wrong b/c of the existence of isotopes
  2. Atoms cannot be divided into smaller particles – wrong b/c of the existence of electrons, protons and neutrons
2. a. Which scientist did the gold foil experiment? Rutherford  
 b. What did he conclude from the gold foil experiment?
  1. Atoms are made up of mostly empty space
  2. Atoms have a small positive center call the nucleus
3. a. Who discovered the electron? J. J. Thomson  
 b. What was the name of his experiment? Cathode Ray Tube Experiment  
 c. What was his model of the atom called? Plum Pudding Model
4. What did Mendeleev do?  
Arranged the periodic table in order of increasing atomic mass
5. a. Who arranged the **modern** periodic table? Moseley  
 b. How is the modern periodic table arranged? Arranged in order of Atomic Number
6. a. What are the three subatomic particles? Give the symbols, charge, mass and location of each.

Particles	Symbol	Charge	Mass	Location
Proton	$p^+$	Positive	1 amu	Nucleus
Neutron	$n^0$	Neutral	1 amu	Nucleus
Electron	$e^-$	Negative	1/2000 amu	Electron Cloud

7. What is special about the proton? **The proton IDENTIFIES THE ELEMENT**
8. Fill in the blanks in the following:
  - a. Mass # =  $p^+ + n^0$
  - b. Atomic # = # of protons
9. What is the difference between Average Atomic mass and mass #? (Average) Atomic Mass is found on the periodic table and is the weighted average of the isotopes. The Mass Number is the mass of one element and must be given or calculated by adding the protons and neutron of that element.
10. Which of the following are isotopes of the same element? How do you know? A, and D. They have the same Atomic Number (# of Protons)
  - a.  $^{12}_5X$
  - b.  $^{15}_5X$
  - c.  $^{12}_6X$
  - d.  $^{15}_7X$
11. Calculate the average atomic mass for Boron if it has two naturally occurring isotopes: boron -10 (abundance of 19.8%) and boron - 11 (abundance of 80.2%).  

$$\{ 10 \times \frac{19.8}{100} \} + \{ 11 \times \frac{80.2}{100} \} = 10.802 \text{ amu}$$
12. Identify an element "X" by calculating the average atomic mass. X = 6 (abundance of 7.5%) and X= 7 (abundance 92.5%).  

$$\{ 6 \times \frac{7.5}{100} \} + \{ 7 \times \frac{92.5}{100} \} = 6.925 \text{ amu} \quad \text{Which is Lithium}$$

# Guide Nuclear Chemistry Answers

**Lingsheng Yao**



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## **Table of Contents Guide Nuclear Chemistry Answers**

1. Understanding the eBook Guide Nuclear Chemistry Answers
  - The Rise of Digital Reading Guide Nuclear Chemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Guide Nuclear Chemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Guide Nuclear Chemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guide Nuclear Chemistry Answers
  - Personalized Recommendations

- Guide Nuclear Chemistry Answers User Reviews and Ratings
- Guide Nuclear Chemistry Answers and Bestseller Lists
- 5. Accessing Guide Nuclear Chemistry Answers Free and Paid eBooks
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  - Guide Nuclear Chemistry Answers eBook Subscription Services
  - Guide Nuclear Chemistry Answers Budget-Friendly Options
- 6. Navigating Guide Nuclear Chemistry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Guide Nuclear Chemistry Answers Compatibility with Devices
  - Guide Nuclear Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Guide Nuclear Chemistry Answers
  - Highlighting and Note-Taking Guide Nuclear Chemistry Answers
  - Interactive Elements Guide Nuclear Chemistry Answers
- 8. Staying Engaged with Guide Nuclear Chemistry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guide Nuclear Chemistry Answers
- 9. Balancing eBooks and Physical Books Guide Nuclear Chemistry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guide Nuclear Chemistry Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Guide Nuclear Chemistry Answers
  - Setting Reading Goals Guide Nuclear Chemistry Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guide Nuclear Chemistry Answers
  - Fact-Checking eBook Content of Guide Nuclear Chemistry Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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