In this activity, we will use a representation of the atom in which a central nucleus containing the protons and neutrons is surrounded by circles that represent the energy levels electrons can occupy.

Draw an atom including a nucleus and five energy levels that electrons could occupy.
 Use a dot to represent an electron at the lowest energy level.



One way an atom emits light (photons) occurs when an electron drops down from a high energy level (also referred to as an excited state) to a lower energy level (the lowest energy level is referred to as the ground state.)

Will an atom emit light if all of the atom's electrons are in the ground state? Explain your reasoning.

No, if you are in the ground state, that is the lowest energy state and you cannot go any lower becasue there is no energy to give up.

- 3) In which case does an atom emit more energy (circle one)?
 - Case A: An electron drops down from the first excited state to the ground state.

 Case B: An electron drops down from the third excited state to the ground state.

Explain your reasoning.

Each excited state has more and more energy, coming down from a higher energy state will release more energy.

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- 1. Understanding the eBook Light And Atoms Lecture Tutorial For Astronomy Ebooks
 - The Rise of Digital Reading Light And Atoms Lecture Tutorial For Astronomy Ebooks
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Light And Atoms Lecture Tutorial For Astronomy Ebooks
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Light And Atoms Lecture Tutorial For Astronomy Ebooks
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Light And Atoms Lecture Tutorial For Astronomy Ebooks

- Personalized Recommendations
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- Light And Atoms Lecture Tutorial For Astronomy Ebooks and Bestseller Lists
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 - Interactive Elements Light And Atoms Lecture Tutorial For Astronomy Ebooks
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 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Light And Atoms Lecture Tutorial For Astronomy Ebooks
- 9. Balancing eBooks and Physical Books Light And Atoms Lecture Tutorial For Astronomy Ebooks
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Light And Atoms Lecture Tutorial For Astronomy Ebooks
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Light And Atoms Lecture Tutorial For Astronomy Ebooks
 - Setting Reading Goals Light And Atoms Lecture Tutorial For Astronomy Ebooks
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Light And Atoms Lecture Tutorial For Astronomy Ebooks

- Fact-Checking eBook Content of Light And Atoms Lecture Tutorial For Astronomy Ebooks
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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