

## Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
  - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
  - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
  - $\frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} \cdot \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} = \frac{2 \text{ mol CaO}}{2 \text{ mol CaO}} = \frac{1 \text{ mol O}_2}{1 \text{ mol CaO}} \cdot \frac{2 \text{ mol CaO}}{1 \text{ mol O}_2}$ . Remember the ratios can be reduced as well!!!
  - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}}$
- $\frac{165.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{165.99 \text{ g Na}_2\text{CO}_3} \cdot \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} = \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \cdot \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} \cdot \frac{49.00 \text{ g NaOH}}{1 \text{ mol NaOH}} = \frac{49.00 \text{ g NaOH}}{1 \text{ mol NaOH}} \cdot \frac{1 \text{ mol NaOH}}{49.00 \text{ g NaOH}} \cdot \frac{108.99 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} = \frac{108.99 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} \cdot \frac{1 \text{ mol CaCO}_3}{108.99 \text{ g CaCO}_3}$
- Molar Mass
  - Definition: the mass of one mole of a substance
  - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen  $\rightarrow$  Water
  - 5.0 mol H<sub>2</sub>
  - 2.5 mol O<sub>2</sub>
- Ethane combustion reaction
  - 15.8 mol O<sub>2</sub>
  - 9.00 mol CO<sub>2</sub> & 13.5 mol H<sub>2</sub>O
- 575 g Na & 886 g Cl<sub>2</sub>
- Iron production reaction
  - 75.1 mol CO
  - 50.1 mol Fe & 75.1 mol CO<sub>2</sub>
- 87,390 g CO or  $8.739 \times 10^5 \text{ g CO}$  & 12,610 g H<sub>2</sub> or  $1.261 \times 10^4 \text{ g H}_2$
- Lightning reaction of N<sub>2</sub> & O<sub>2</sub>
  - 1,100 g NO<sub>2</sub> or  $1.10 \times 10^3 \text{ g NO}_2$
  720. g NO or  $7.20 \times 10^2 \text{ g NO}$
- CO<sub>2</sub> removal in spacecraft
  - 42.04 mol NaOH
  - 21.02 mol Na<sub>2</sub>CO<sub>3</sub> & 21.02 mol H<sub>2</sub>O
- Double replacement AgNO<sub>3</sub> & NaBr
  - 463 g NaBr
  - 845 g AgBr
- Soda-acid fire extinguisher
  - 3.058 mol NaHCO<sub>3</sub>
  - 3.058 mol CO<sub>2</sub>, 1.529 mol Na<sub>2</sub>SO<sub>4</sub>, 3.058 mol H<sub>2</sub>O
- Sulfuric acid-sodium hydroxide neutralization reaction
  - H<sub>2</sub>SO<sub>4</sub> + 2 NaOH  $\rightarrow$  Na<sub>2</sub>SO<sub>4</sub> + 2 H<sub>2</sub>O
  - 37 g H<sub>2</sub>SO<sub>4</sub>
  - 53 g Na<sub>2</sub>SO<sub>4</sub> & 14 g H<sub>2</sub>O
- Copper & silver nitrate reaction
  - 0.0104 mol Cu(NO<sub>3</sub>)<sub>2</sub>
  - 0.0104 mol Cu & 0.0209 mol AgNO<sub>3</sub>
- Aspirin production
  - 13.5 kg C<sub>6</sub>H<sub>5</sub>O<sub>4</sub>
  - 7.66 kg C<sub>6</sub>H<sub>5</sub>O<sub>3</sub>
  - 4.29 L HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

# Guided Reading And Study Workbook Chapter 9

## Stoichiometry Answers

**Prentice Hall PTR**



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### **Table of Contents Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers**

1. Understanding the eBook Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
  - The Rise of Digital Reading Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers

- Personalized Recommendations
  - Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers User Reviews and Ratings
  - Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers and Bestseller Lists
5. Accessing Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers Free and Paid eBooks
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  - Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers eBook Subscription Services
  - Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers Budget-Friendly Options
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  - Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
  - Highlighting and Note-Taking Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
  - Interactive Elements Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
8. Staying Engaged with Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
9. Balancing eBooks and Physical Books Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
- Setting Reading Goals Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers

- Fact-Checking eBook Content of Guided Reading And Study Workbook Chapter 9 Stoichiometry Answers
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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tsa 5 ab lsa here a 4 81 cm lsa 105 cm 2

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