

Acids and bases

Chapter outline

- Recall the definitions of Brønsted–Lowry acids and bases and deduce these in a chemical reaction, including the identification of conjugate acid–base pairs.
- Recall the terms ‘amphoteric’ and ‘amphiprotic’.
- Describe the reactions of acids with metals, metal oxides, metal hydroxides, hydrogen carbonates and carbonates.
- Recall that pH is a measure of the concentration of aqueous hydrogen ions and solve problems involving pH and the concentration of hydrogen ions and hydroxide ions.
- Explain the difference between strong and weak acids and bases and describe experiments that could be used to distinguish these.
- Recall what is meant by acid deposition and explain some of the problems associated with acid deposition as well as pre- and post-combustion methods of reducing sulfur oxide emissions.
- Recall the definitions of Lewis acids and bases and recognise these in a chemical reaction. **HL**
- Deduce the dissociation constant for weak acids and bases. **HL**
- Sketch and explain the important features of titration curves including the composition and action of a buffer and use these curves to deduce the pK_a or pK_b value of a weak acid or base. **HL**
- Explain how an acid–base indicator works and select a suitable indicator for a titration. **HL**

KEY TERMS

Acid deposition: A more general term than acid rain; it refers to any process in which acidic substances leave the atmosphere and are deposited on the Earth’s surface.

Brønsted–Lowry acid: A proton (H^+) donor.

Brønsted–Lowry base: A proton (H^+) acceptor.

pH: A measure of the concentration of hydrogen ions in an aqueous solution.
 $pH = -\log_{10}[H^+]$

Strong acid/base: An acid/base that is completely dissociated into its ions in aqueous solution.

Weak acid/base: An acid/base that is only partially dissociated into its ions in aqueous solution.

Guided Study Workbook Acids And Bases Answers

Neil E. Schore



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

Study Guide and Solutions Manual Neil E. Schore, 2002-08-02 *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Organic Chemistry, 13e Student Study Guide and Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2022-05-03 Organic Chemistry Student Study Guide and Solutions Manual 13th Edition offers the full solutions for select exercises from the text

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chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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Study Guide with Answers to Selected Problems H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

Organic Chemistry, Study Guide and Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, 2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

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Organic Chemistry Michael B. Smith, 2016-03-09 Based on the premise that many if not most reactions in organic chemistry can be explained by variations of fundamental acid base concepts Organic Chemistry An Acid Base Approach provides a framework for understanding the subject that goes beyond mere memorization Using several techniques to develop a relational understanding it helps students fully grasp the essential concepts at the root of organic chemistry This new edition was rewritten largely with the feedback of students in mind and is also based on the author's classroom experiences using the first edition Highlights of the Second Edition Include Reorganized chapters that improve the presentation of material Coverage of new topics such as green chemistry Adding photographs to the lectures to illustrate and

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