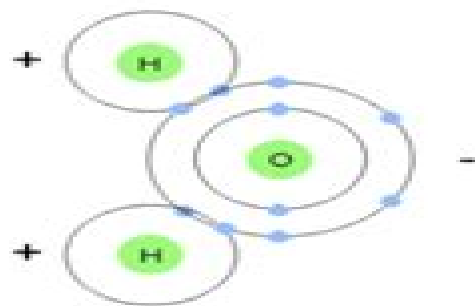


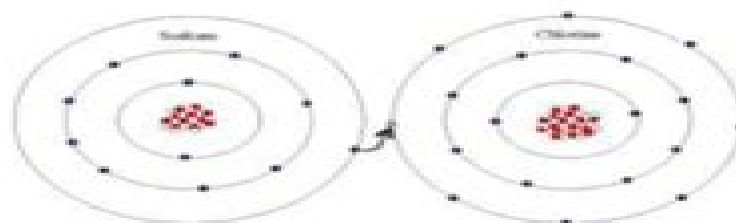
## Chemical Bonding (PS 4c)

Test date: \_\_\_\_\_

Essential Concepts and Skills	Assignments
<p><b>The Formation of Compounds through Ionic and Covalent Bonding (PS 4c)</b></p> <ul style="list-style-type: none"> <li>Metals tend to lose electrons in chemical reactions, forming positive ions. Nonmetals tend to gain electrons in chemical reactions, forming negative ions.</li> <li>Given a chemical formula of a compound, identify the elements and the number of atoms of each that comprise the compound.</li> <li>Recognize that the number of electrons on the outermost energy level (valence electrons) determines an element's chemical properties or chemical reactivity.</li> <li>Describe the difference between ionic and covalent bonding.</li> <li>Predict what kind of bond (ionic or covalent) will likely form when metals and nonmetals are chemically bonded.</li> <li>Gaining or losing electrons makes an atom an ion.</li> <li>Atoms react to form chemically stable substances that are held together by chemical bonds and are represented by chemical formulas. To become chemically stable, atoms gain, lose, or share electrons.</li> <li>Compounds are formed when elements chemically react. When a metallic element reacts with a nonmetallic element, their atoms gain and lose electrons respectively, forming ionic bonds. Generally, when two nonmetals react, atoms share electrons, forming covalent (molecular) bonds.</li> </ul>	<p>Pages 328-345 Starter Quiz 1, 2 Gizmos</p>



**Covalent Bond**



**Ionic Bond**

# Ionic And Covalent Bonding Study Guide

**IM Harris**



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