

Ionic and Covalent Bonds Study Guide

Be able to define the following terms:

- | | | |
|-------------------|-----------------|--------------------|
| ➤ covalent bond | ➤ anion | ➤ polyatomic ion |
| ➤ Lewis structure | ➤ cation | ➤ hydrogen bonding |
| ➤ molecule | ➤ chemical bond | ➤ valence electron |
| ➤ polar bond | ➤ ionic bond | |
| ➤ non-polar bond | ➤ ion | |

1. A group of covalently bonded atoms that acts together as one charged atom is a ____.
2. A chemical bond that occurs when atoms share electrons is a(n) ____ bond.
3. The ____ tells you how many electrons an atom must gain, lose, or share to become stable.
4. The oxidation number of an atom is shown with a ____.
5. How many electrons are needed in the outer energy levels of most atoms for the atom to be chemically stable?
6. Why do the noble gases NOT form compounds readily?
7. What is the total number of atoms in the compound $\text{Ca}(\text{ClO}_3)_2$?
8. Explain why it is correct to say that a compound like calcium carbonate, CaCO_3 , contains both ionic and covalent bonds.
9. Name or write the formulas for the following ionic compounds:
 - a. Al_2O_3
 - b. NH_4Cl
 - c. FeO
 - d. manganese (V) oxide
 - e. calcium chloride
 - f. lead (IV) nitrate
10. Name or write the formulas for the following covalent compounds
 - a. P_2O_5
 - b. HF
 - c. HClO
 - d. carbon tetrachloride
 - e. hydrochloric acid
 - f. carbonic acid

Ionic And Covalent Bonds Study Guide

Tyler L. Jackson



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Know Your 'O' Level Chemistry - A Study Guide , **Basic Concepts of Chemistry, 9e Study Guide and Solutions**

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coauthor Neil Schore this invaluable manual includes chapter introductions that highlight new materials chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

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are included in the Appendix Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

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Table of Contents Ionic And Covalent Bonds Study Guide

1. Understanding the eBook Ionic And Covalent Bonds Study Guide
 - The Rise of Digital Reading Ionic And Covalent Bonds Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic And Covalent Bonds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic And Covalent Bonds Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic And Covalent Bonds Study Guide
 - Personalized Recommendations
 - Ionic And Covalent Bonds Study Guide User Reviews and Ratings
 - Ionic And Covalent Bonds Study Guide and Bestseller Lists
5. Accessing Ionic And Covalent Bonds Study Guide Free and Paid eBooks
 - Ionic And Covalent Bonds Study Guide Public Domain eBooks
 - Ionic And Covalent Bonds Study Guide eBook Subscription Services
 - Ionic And Covalent Bonds Study Guide Budget-Friendly Options
6. Navigating Ionic And Covalent Bonds Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic And Covalent Bonds Study Guide Compatibility with Devices
 - Ionic And Covalent Bonds Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic And Covalent Bonds Study Guide
 - Highlighting and Note-Taking Ionic And Covalent Bonds Study Guide
 - Interactive Elements Ionic And Covalent Bonds Study Guide
8. Staying Engaged with Ionic And Covalent Bonds Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Ionic And Covalent Bonds Study Guide
- 9. Balancing eBooks and Physical Books Ionic And Covalent Bonds Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic And Covalent Bonds Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic And Covalent Bonds Study Guide
 - Setting Reading Goals Ionic And Covalent Bonds Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic And Covalent Bonds Study Guide
 - Fact-Checking eBook Content of Ionic And Covalent Bonds Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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