Find the atomic mass of the following compounds

Calcium oxide (CaO)

Water

Classify (ionic, covalent or polyatomic) and name the following compounds.

- 1. CO<sub>2</sub>
- 3. BeCrO<sub>4</sub> \_\_\_\_\_

Find the valence electrons and oxidation number of:

2. Oxygen 1. Heliann

Draw the Lewis Dot Diagrams for:

L. Carbon.

2. Sodium

Make Balanced Compounds from:

- I. Li + 0
- 2. Na + Ar \_\_\_\_\_
- 3. Be + NO, 1.
- 4. Ca + O

With dot diagrams draw the covalient bond of Oc.

With dot diagrams the covalent bond of OF at

Matching: A. Mistage A substance made up of two or more atoms

Bt. Compound

- C. Element
- D. Matter
- E. Heterogenous
- F. Homogeneous
- that must be separated by chemical means.

Something made up of 2 or more substances that can be separated by physical means.

A substance that is pure and made up of only one type of atom.

- Anything that has mass and takes up space.
- A mixture that is the same throughout.
- A mixture that is different throughout.

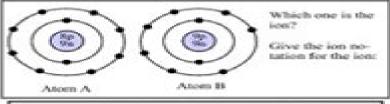
1. Valence electrons are the outermost electrons of an atom that are involved in chemical bonding? True/False.

2. Lithium has \_\_\_\_\_\_ valence electrons, will (lose or gain) electrons and become (positive or negative). This is why lithium's oxidation number is

 A metal and non-metal will form a compound; two non-metals form a \_\_\_\_\_\_ compound.

4. Protons are \_\_\_\_; electrons are \_\_\_\_; neu-

5. Protons and neutrons are in the center of the atom, which is called the



Give the number of protons for:

- L. Carbon \_\_\_\_\_\_ 3. Iron \_\_\_\_\_
- 2. Beryllium 4. Chlorine

If you change the number of protons you Martchings change the If you change the number of neutrons, you AL SOM change the B. Element C. Isotope If you change the number of electrons, you DOM: change the Every 5,000 miles you should change a ear's

Marchings A. Law of Com-

- servation of Mass
- Bl. Meter C. Centimeter
- D. Litter
- E. Grunn

Unit of muse; about 1 dollar bill. 1/100th of meter; width of pinky Unit of volume; just b igger a Unit of length; 3.3.

Manghing: A. Solid

District At

C. Cias

IDV: 4DGE

Molecules that are tightly packed and retain their shape and size.

Mass is neither destroyed or created in chemical reactions.

Molecules that bounce off of each other, can be compressed, and take the shape of their container.

Molecules that can move (slide) over each other and have a definite size (volume), but not shape and can not be compressed.

# **Ipc Chemistry Review**

**JR Anderson** 



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