

## Stoichiometry Lab: Iron and Copper Sulfate Sample data

### Data and observations :

mass of copper (II) sulfate pentahydrate	<u>3.27</u>	g
mass of iron:	<u>0.50</u>	g
mass of filter paper	<u>1.28</u>	g
mass of filter paper and dry copper:	<u>1.84</u>	g
mass of dry copper	<u>0.56</u>	g

### Observations of reaction:

When the iron was added a precipitate began to form. The precipitate was reddish in color. The solution remained a blue color but was somewhat lighter in color.

### Calculations:

1&2. The single replacement equation will be:  $\text{Fe} + \text{CuSO}_4 \longrightarrow \text{FeSO}_4 + \text{Cu}$

3. Fe and  $\text{CuSO}_4$  are the reactants and  $\text{FeSO}_4$  and Cu are the products.

4. The moles of copper (II) sulfate pentahydrate used were 0.013 mol.

$$3.27 \text{ g CuSO}_4 \cdot 5\text{H}_2\text{O} \times \frac{1 \text{ mol CuSO}_4 \cdot 5\text{H}_2\text{O}}{249.47 \text{ g CuSO}_4 \cdot 5\text{H}_2\text{O}} = 0.013 \text{ mol CuSO}_4 \cdot 5\text{H}_2\text{O}$$

5. The moles of iron used were 0.009 mol Fe.

$$0.50 \text{ g Fe} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} = 0.009 \text{ mol Fe}$$

6. The moles of copper produced were 0.009 mol Cu.

$$0.56 \text{ g Cu} \times \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} = 0.009 \text{ mol Cu}$$

# Iron And Copper Sulfate Lab Answers

**C Cleary**

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