

Chemical Reactions

Section 9.1 Reactions and Equations

pages 282–288

Practice Problems

pages 284–287

Write skeleton equations for the following word equations.

1. Hydrogen and bromine gases react to yield hydrogen bromide.

hydrogen(g) + bromine(g) →
hydrogen bromide(g)



2. When carbon monoxide and oxygen react, carbon dioxide forms.

carbon monoxide(g) + oxygen(g) →
carbon dioxide(g)



3. **Challenge** Write the word equation and the skeleton equation for the following reaction: when heated, solid potassium chlorate yields solid potassium chloride and oxygen gas.

Potassium chlorate(s) →
potassium chloride(s) + oxygen(g)



Write chemical equations for each of the following reactions.

4. In water, iron(III) chloride reacts with sodium hydroxide, producing solid iron(III) hydroxide and sodium chloride.



5. Liquid carbon disulfide reacts with oxygen gas, producing carbon dioxide gas and sulfur dioxide gas.



6. **Challenge** A piece of zinc metal is added to a solution of hydrogen sulfate. This reaction produces a gas and a solution of zinc sulfate.



Section 9.1 Assessment

page 288

7. **Explain** why it is important that a chemical equation be balanced.

Because mass is neither created nor destroyed in chemical reactions, the numbers of atoms of all elements must be equal on both sides of the reaction arrow.

8. **List** three types of physical evidence that indicate a chemical reaction has occurred.

Answers may include release or absorption of energy, change in color, change in odor, formation of a gas, or formation of a solid.

9. **Compare and contrast** a skeleton equation and a chemical equation.

The skeleton equation includes the formulas of reactants and products. The chemical equation gives the relative amounts of reactants and products.

10. **Explain** why it is important to reduce coefficients in a balanced equation to the lowest possible whole-number ratio.

Coefficients in the lowest ratio most clearly indicate the relative amounts of substances in a reaction.

11. **Analyze** When balancing a chemical equation, can you adjust the subscript in a formula? Explain.

No. Doing so changes the identity of the substance.

Mcgraw Hill Chemical Reactions Study Guide

Krishina Subramanyam



Mcgraw Hill Chemical Reactions Study Guide:

Scientific and Technical Information Resources Krishina Subramanyam,2020-07-24 This book focuses on current practices in scientific and technical communication historical aspects and characteristics and bibliographic control of various forms of scientific and technical literature It integrates the inventory approach for scientific and technical communication

Chemical Processing Handbook John J. McKetta Jr,1993-04-30 Written by more than 40 world renowned authorities in the field this reference presents information on plant design significant chemical reactions and processing operations in industrial use offering shortcut calculation methods wherever possible **Glencoe Science** McGraw-Hill Staff,2001-06

Library of Congress Catalog: Motion Pictures and Filmstrips Library of Congress,1968 *Macmillan/McGraw-Hill Math: Teacher ed., v. 2*,2004 **Ludwig's Applied Process Design for Chemical and Petrochemical Plants** A. Kayode Coker,2014-11-29 The fourth edition of Ludwig s Applied Process Design for Chemical and Petrochemical Plants Volume Three is a core reference for chemical plant and process engineers and provides an unrivalled reference on methods process fundamentals and supporting design data New to this edition are expanded chapters on heat transfer plus additional chapters focused on the design of shell and tube heat exchangers double pipe heat exchangers and air coolers Heat tracer requirements for pipelines and heat loss from insulated pipelines are covered in this new edition along with batch heating and cooling of process fluids process integration and industrial reactors The book also looks at the troubleshooting of process equipment and corrosion and metallurgy Assists engineers in rapidly analyzing problems and finding effective design methods and mechanical specifications Definitive guide to the selection and design of various equipment types including heat exchanger sizing and compressor sizing with established design codes Batch heating and cooling of process fluids supported by Excel programs *A Study of Enzymes* Stephen A. Kuby,2019-07-23 First published in 1990 this comprehensive monograph consists of two parts Volume I entitled Enzyme Catalysis Kinetics and Substrate Binding and Volume II entitled Mechanism of Enzyme Action Volume I focuses on several aspects of enzyme catalytic behavior their steady state and transient state kinetics and the thermodynamic properties of substrate binding Packed with figures tables schemes and photographs this volume contains over 1 000 references including references regarding enzymology s fascinating history This comprehensive book is of particular interest to enzymology students teachers and researchers Volume II presents selected cutting edge examples of techniques and approaches being pursued in biochemistry This up to date resource includes 11 chapters which illustrate important theoretical and practical aspects of enzyme mechanisms It also features selected examples in which today s most important techniques ideas and theories are used to elaborate on the intricate nature of enzyme action mechanisms This particular volume provides important information for both the novice and the seasoned investigator U.S. Environmental Protection Agency Library System Book Catalog Holdings as of July 1973 United States. Environmental Protection Agency. Library Systems Branch,1974 *Advances in Theory and Practice of Computational*

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