

## **AP Biology Study Guide Unit 1**

### **Properties of Water:**

1. Polarity - allows H<sub>2</sub>O to form Hydrogen Bonds
2. Cohesion - Water H bonds to itself, creating surface tension
3. Adhesion - Water adheres to other polar substances
  - a. Allows easy transport of water against Gravity
4. Universal Solvent - Water can dissolve Polar and Ionic Compounds
  - a. This allows for reactions between Aqueous Solutions
5. Expands when Frozen - this means that Ice floats
6. High Specific Heat - water can absorb or release a large amount of heat with only a slight change in its own temperature (caused by cohesion)
  - a. High Heat of Vaporization - only warmest molecules can evaporate, so this causes evaporative cooling
7. Can be regulated by non-polar hydrophobic substances
8. Neutral PH → can form acids or bases if substances are added

Polarity - Difference in electronegativity between atoms bonded together create slightly charged ends of the molecules they form

Hydrophobic Substance - Nonpolar substance that repels water

Hydrophilic Substance - Polar substance that attracts water molecules

### **CHON - Carbon, Hydrogen, Nitrogen, Oxygen**

- Makeup 99% of living matter
- Strong Biological Cornerstones Because:
  - Smallest Diameter/Valency for atoms of 1, 2, 3, 4 valence electrons
  - Form Stable Covalent Bonds
  - Can form double and triple bonds
  - Versatile

### **Biological Transition Metals - Mn, Fe, Co, Cu, Zn**

- Can exist in multiple oxidation states → can be both oxidized and reduced

### **Biological Ions: Na<sup>+</sup>, Mg<sup>2+</sup>, Cl<sup>-</sup>, K<sup>+</sup>, Ca<sup>2+</sup>**

- Used to maintain homeostasis and osmotic concentration

Phosphorous/Sulphur - Energy Carriers which form strong high energy bonds

### **5 Foundation Biological Molecules:**

# Mcgraw Hill Ap Biology Study Guide

**Mark Anestis, Kellie Ploeger Cox**



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## **Table of Contents Mcgraw Hill Ap Biology Study Guide**

1. Understanding the eBook Mcgraw Hill Ap Biology Study Guide
  - The Rise of Digital Reading Mcgraw Hill Ap Biology Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Ap Biology Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mcgraw Hill Ap Biology Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Ap Biology Study Guide
  - Personalized Recommendations

- Mcgraw Hill Ap Biology Study Guide User Reviews and Ratings
- Mcgraw Hill Ap Biology Study Guide and Bestseller Lists
- 5. Accessing Mcgraw Hill Ap Biology Study Guide Free and Paid eBooks
  - Mcgraw Hill Ap Biology Study Guide Public Domain eBooks
  - Mcgraw Hill Ap Biology Study Guide eBook Subscription Services
  - Mcgraw Hill Ap Biology Study Guide Budget-Friendly Options
- 6. Navigating Mcgraw Hill Ap Biology Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Mcgraw Hill Ap Biology Study Guide Compatibility with Devices
  - Mcgraw Hill Ap Biology Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mcgraw Hill Ap Biology Study Guide
  - Highlighting and Note-Taking Mcgraw Hill Ap Biology Study Guide
  - Interactive Elements Mcgraw Hill Ap Biology Study Guide
- 8. Staying Engaged with Mcgraw Hill Ap Biology Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mcgraw Hill Ap Biology Study Guide
- 9. Balancing eBooks and Physical Books Mcgraw Hill Ap Biology Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mcgraw Hill Ap Biology Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mcgraw Hill Ap Biology Study Guide
  - Setting Reading Goals Mcgraw Hill Ap Biology Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcgraw Hill Ap Biology Study Guide
  - Fact-Checking eBook Content of Mcgraw Hill Ap Biology Study Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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