

HEAT & TEMPERATURE



READINGS
ACTIVITIES
ASSESSMENTS

ALBERTA — GRADE 7 — SCIENCE

Three sample worksheets are displayed, overlapping each other. The worksheets are titled "Heat vs Temperature", "Boiling Cuts - Sidney's Test", and "Heat vs Temperature".

The "Heat vs Temperature" worksheet includes a table with columns for Time and Temperature, and a section for Questions.

	Time	Temperature
1)	1:00 pm	
2)	1:05 pm	
3)	1:10 pm	
4)	1:15 pm	
5)	1:20 pm	
6)	1:25 pm	
7)	1:30 pm	
8)	1:35 pm	
9)	1:40 pm	
10)	1:45 pm	

The "Boiling Cuts - Sidney's Test" worksheet includes a table with columns for Time and Temperature, and a section for Questions.

	Time	Temperature
1)	1:00 pm	
2)	1:05 pm	
3)	1:10 pm	
4)	1:15 pm	
5)	1:20 pm	
6)	1:25 pm	
7)	1:30 pm	
8)	1:35 pm	
9)	1:40 pm	
10)	1:45 pm	

The "Heat vs Temperature" worksheet includes a section for Questions and a diagram of a pot of water on a stove.

Heat vs Temperature

Heat is a form of energy that is transferred from one object to another. It is measured in Joules (J). Temperature is a measure of the average kinetic energy of the particles in a substance. It is measured in degrees Celsius (°C).

Heat and temperature are related, but they are not the same thing. Heat is the energy that is transferred, while temperature is the measure of the average kinetic energy of the particles.

When you heat a substance, you are adding energy to it. This causes the particles to move faster, which increases the temperature. However, the temperature will only increase if the substance is not changing state. For example, when you heat ice, it will melt into water, but the temperature will stay at 0°C until all the ice has melted.

Similarly, when you cool a substance, you are removing energy from it. This causes the particles to move slower, which decreases the temperature. However, the temperature will only decrease if the substance is not changing state. For example, when you cool water, it will freeze into ice, but the temperature will stay at 0°C until all the water has frozen.

Heat and temperature are both important concepts in science. Understanding the difference between them is essential for many experiments and calculations.

Key Differences

- Heat is a form of energy, while temperature is a measure of the average kinetic energy of the particles.
- Heat is measured in Joules (J), while temperature is measured in degrees Celsius (°C).
- Heat can be transferred from one object to another, while temperature cannot.
- Heat is a process, while temperature is a state.

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**Louise Finlaison, Edmonton Public
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