

- Saponifiable
- Saturated solution
- Solid
- Solubility
- Solute
- Solution
- Solvent
- Specific heat

- Standard temperature and pressure
- Stoichiometry
- Sublimation
- Substrate
- Supersaturated solution
- Surface tension
- Temperature
- Theoretical yield

- Thermochemical equation
- Thermochemistry
- Titration
- Triprotic acid
- Unsaturated solution
- Vapor pressure lowering
- Vaporization
- Viscosity

CONCEPT QUESTIONS: You should be able to answer/describe each of the following in order to gain a real understanding of the different concepts we have covered this second semester. You can expect to see a majority of the questions on the exam similar to the ones below.

Unit 5 (Chapters 6 & 7)

1. Convert each of the following to either moles or particles.

a. 6.022×10^{23} atoms He

$$6.022 \times 10^{23} \text{ atoms He} \times \frac{1 \text{ mol He}}{6.022 \times 10^{23} \text{ atoms He}} = \boxed{1.00 \text{ mol He}}$$

b. 2.56 mol Ne

$$2.56 \text{ mol Ne} \times \frac{6.022 \times 10^{23} \text{ atoms Ne}}{1 \text{ mol Ne}} = \boxed{1.54 \times 10^{24} \text{ atoms Ne}}$$

2. Find the molar mass for each of the following compounds.

a. CHCl_3

$$12.01 + 1.01 + 35.45(3) = \boxed{119.37 \text{ g/mol}}$$

b. HMnO_4

$$1.01 + 54.94 + 16.00(4) = \boxed{119.95 \text{ g/mol}}$$

3. Convert each of the following to either moles or mass.

a. 16.35 g PbO

$$16.35 \text{ g PbO} \times \frac{1 \text{ mol PbO}}{223.20 \text{ g PbO}} = \boxed{.07 \text{ mol PbO}}$$

b. 0.025 mol Na_2SO_4

$$.025 \text{ mol Na}_2\text{SO}_4 \times \frac{142.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = \boxed{3.55 \text{ g Na}_2\text{SO}_4}$$

4. Convert each of the following to moles or volume.

a. 22.4 mol H_2

$$22.4 \text{ mol H}_2 \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = \boxed{501.76 \text{ L H}_2}$$

b. 0.145 L F_2

$$.145 \text{ L F}_2 \times \frac{1 \text{ mol F}_2}{22.4 \text{ L F}_2} = \boxed{.01 \text{ mol F}_2}$$

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PT Brinkman



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