

Installation Manual of Room Air Conditioner

Preparation

Necessary Tools for Installation

- Hammer
- Nipper
- Hack saw
- Hole core drill
- Spanner(17,19 and 25mm)
- Gas leakage detector or soap-and-water solution
- Torque wrench (17mm,22mm,26mm)
- Pipe cutter
- Flaring tool
- Knife
- Measuring tape
- Router

Power Source

- Before inserting power plug into receptacle, check the voltage without fail. The power source is the same as the corresponding name plate.
- Install an exclusive branch circuit of the power.
- A receptacle shall be set up in a distance where the power cable can be reached. Do not extend the cable by cutting it.

Selection of Installation Place

Indoor Unit - Select a location that is

- Robust not causing vibration, where the body can be supported sufficiently.
- Not affected by heat or steam generated in the vicinity, where inlet and outlet of the unit are not disturbed.
- Possible to drain easily, where piping can be connected with the outdoor unit.
- Where cold air can be spread in a room evenly.
- Nearby a power receptacle. (Refer to drawings).
- Place where the distance of more than 1m from televisions, radios, wireless apparatuses and fluorescent lamps can be left.
- In the case of fixing the remote controller on a wall, place where the indoor unit can receive signals when the fluorescent lamps in the room are in use.

Outdoor Unit - Select a location that is

- Less affected by rain or direct sunlight and is sufficiently ventilated.
- Strong enough to bear the unit, where vibration and noise are not increased.
- Not causing a nuisance to neighbors due to discharged air or noise.
- A distance marked as is available as illustrated in the below figure.

Drawing for the installation of indoor and outdoor units

The models adopt HCFC free refrigerant R410A.



The above picture is for your reference only. Your product may look different. Read this manual before installation. Explain the operation of the unit to the user according to this manual.

NO.0010536242

Hino Bus Air Conditioning Manual Hand

Rachel S Tattersall



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jan 3 2023 a net ionic equation is a chemical equation that shows only the species participating in a chemical reaction and omits the spectator ions spectator ions are ions that

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations

requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction

therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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$1\text{K}_3\text{PO}_4 + 3\text{Al}(\text{OH})_3 \rightarrow 3\text{KOH} + 3\text{AlPO}_4$ ionic $3\text{K}^+ + \text{PO}_4^{3-} + 3\text{Al}(\text{OH})_3 \rightarrow 3\text{K}^+ + 3\text{OH}^- + 3\text{AlPO}_4$ net ionic $\text{PO}_4^{3-} + 3\text{Al}(\text{OH})_3 \rightarrow 3\text{OH}^- + 3\text{AlPO}_4$

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds cl aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh₄ aq cr₂o₇²⁻ aq ag₂cr₂o₇ s 2nh₄ aq 2f aq because two nh₄ aq

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