



A thermometer is secured with a thermometer clamp to guard against breakage.

# Experiment 14

## Molar Mass of a Solid

- To observe and measure the effect of a solute on the freezing point of a solvent
- To determine the molar mass (molecular weight) of a nonvolatile, nonelectrolyte solute

### OBJECTIVES

The following techniques are used in the Experimental Procedure:



### TECHNIQUES

A pure liquid, such as water or ethanol, has characteristic physical properties: the melting point, boiling point, density, vapor pressure, viscosity, surface tension, and additional data listed in handbooks of chemistry. The addition of a **soluble** solute to the liquid forms a homogeneous mixture called a **solution**. The solvent of the solution assumes physical properties that are no longer definite but dependent on the amount of solute added. The vapor pressure of the solvent decreases, the freezing point of the solvent decreases, the boiling point of the solvent increases, and the osmotic pressure of the solvent increases. The degree of the change depends on the *number* of solute particles that have dissolved, *not* on the chemical identity of the solute. These four physical properties that depend on the number of solute particles dissolved in a solvent are called **colligative properties**.

For example, one mole of glucose or urea (neither of which dissociates in water) lowers the freezing point of one kilogram of water by  $1.86^{\circ}\text{C}$ ; whereas one mole of sodium chloride lowers the freezing point of one kilogram of water by nearly twice that amount ( $\sim 3.72^{\circ}\text{C}$ ) because, when dissolved in water, it dissociates into  $\text{Na}^{+}$  and  $\text{Cl}^{-}$  providing twice as many moles of solute particles per mole of solute as do glucose or urea.

When freezing ice cream at home, a salt-ice water mixture provides a lower temperature bath than an ice/water mixture alone. Antifreeze (ethylene glycol, Figure 14.1) added to the cooling system of an automobile reduces the probability of freeze-up in the winter and boiling over in the summer because the antifreeze/water solution has a lower freezing point and a higher boiling point than pure water.

These changes in the properties of pure water that result from the presence of a **nonvolatile solute** are portrayed by the phase diagram in Figure 14.2, page 190, a plot of vapor pressure versus temperature. The solid lines refer to the equilibrium conditions between the respective phases for pure water; the dashed lines represent the same conditions for an aqueous solution.

### INTRODUCTION



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**Figure 14.1** Ethylene glycol is a major component of most antifreeze solutions.

**Colligative properties:** properties of a solvent that result from the presence of the number of solute particles in the solution and not their chemical composition

**Nonvolatile solute:** a solute that does not have a measurable vapor pressure

# J A Beran Lab Manual Experiment 14

**Beran**



## J A Beran Lab Manual Experiment 14:

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