

## Section 2–1 The Nature of Matter (pages 35–39)

This section identifies the three particles that make up atoms. It also explains how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds.

### Atoms (page 35)

1. The basic unit of matter is called a(an) .
2. Describe the nucleus of an atom.
3. Complete the table about subatomic particles.

| SUBATOMIC PARTICLES |          |                  |
|---------------------|----------|------------------|
| Particle            | Charge   | Location in Atom |
|                     | Positive |                  |
|                     | Neutral  |                  |
|                     | Negative |                  |

4. Why are atoms neutral despite having charged particles?

### Elements (page 36)

5. What is a chemical element?
6. What does an element's atomic number represent?

### Chemical Compounds (page 37)

7. What is a chemical compound?
8. What are the elements in salt? How many of each element are there?

### Chemical Bonds (pages 38–39)

9. What holds atoms in compounds together?
10. Complete the table about the main types of chemical bonds.

## CHEMICAL BONDS

| Chemical bonds |                |
|----------------|----------------|
| Type           | Formed When... |
| Covalent bond  |                |
| Ionic bond     |                |

11. What is an ion?
12. Is the following sentence **true or false**?  
An atom that loses electrons has a negative charge.
13. The structure that results when atoms are joined together by covalent bonds is called a(an) .

# Guided Reading And Study Workbook Chapter 2 Answers Biology

**J Dewey**



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