










Isotopes

Are all atoms of an element alike?

Why?

The following activity will help you learn the important structural characteristics of an atom. How do we classify atoms? How does the combination of subatomic particles affect the mass and charge of an atom? What are isotopes? This is just a sampling of what we will address. Throughout this activity you will want to keep both Model 1 and a periodic table handy.

Model 1

Isotopes of Hydrogen			
Symbol	${}^1_1\text{H}$	${}^2_1\text{H}$	${}^3_1\text{H}$
Atomic Diagram with Name	 Hydrogen-1 (protium)	 Hydrogen-2 (deuterium)	 Hydrogen-3 (tritium)
Number of Protons	1	1	1
Number of Neutrons	0	1	2
Isotopes of Carbon			
Symbol	${}^{12}_6\text{C}$	${}^{13}_6\text{C}$	${}^{14}_6\text{C}$
Atomic Diagram with Name	 Carbon-12	 Carbon-13	 Carbon-14
Number of Protons	6	6	6
Number of Neutrons	6	7	8
Isotopes of Magnesium			
Symbol	${}^{24}_{12}\text{Mg}$	${}^{25}_{12}\text{Mg}$	${}^{26}_{12}\text{Mg}$
Atomic Diagram with Name	 Magnesium-24	 Magnesium-25	 Magnesium-26
Number of Protons	12	12	12
Number of Neutrons	12	13	14

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Isotopes Answers Chemistry Pogil

CL Gary



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