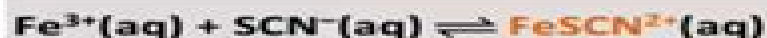


IA lab work on spectrophotometric determination of K_c of the Iron(3)Thiocyanate system



$$K = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]}$$

Since I will not change the temperature in my experiment, the equilibrium constant will stay the same, and therefore I can measure the light absorbance and use the Lambert Beer-Law to find the initial concentration of each species in the system and then calculate the K_c .

$$X(\text{initial concentration of Iron3 and Thiocyanate}) = A(\text{absorbance of Iron(3)Thiocyanate solution}) / \epsilon \cdot l$$

Equipment:

- 4 test tubes with 4 corks
- test tube rack
- 10 ml pipette with pump thing

Solutions:

- 20 ml of Iron (3) solution
- 20 ml of Thiocyanate solution
- 20 ml of water H_2O

Samples	The amount of Thiocyanate solution SCN^{-}	The amount of water H_2O	Iron 3 solution Fe^{3+}
1	2 ml	8 ml	5 ml
2	4 ml	6 ml	5 ml
3	5 ml	5 ml	5 ml
4	9 ml	1 ml	5 ml

Kc Of Iron Thiocyanate Lab Report

K Morrison

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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