

# ACIDS, BASES AND SALTS

<b>Definition of an acid</b> An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	<b>What causes the acidic properties of acids?</b> The hydrogen ions (H <sup>+</sup> ) cause the acidic properties and these are formed in the presence of water.
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<b>Another term to refer to an acid</b> An acid is called a <u>proton donor</u>	<b>Why an acid is also called a proton donor?</b> It's because an acid provides protons or hydrogen ions (H <sup>+</sup> ) to other substances during the reaction.
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<b>Substance to which an acid provides protons</b> BASE	<b>A base is then called</b> PROTON ACCEPTOR	<b>Why the base is called a proton acceptor?</b> Because the base accepts hydrogen ions from acids	<b>Equation for the reaction between acid and base</b> $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$
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<b>Common laboratory acids</b>	<b>These three common laboratory acids are also called</b>	<b>Why these acids are also called mineral acids?</b>
Hydrochloric acid Sulphuric acid Nitric acid	HCl H <sub>2</sub> SO <sub>4</sub> HNO <sub>3</sub>  MINERAL ACIDS	They are derived from mineral salts i.e chlorides for HCl, sulphates for H <sub>2</sub> SO <sub>4</sub> and nitrates for HNO <sub>3</sub>

<b>Other mineral acids known</b>	<b>Mineral salts from which the acid is derived</b>	<b>Organic acids known</b>	<b>Naturally occurring acids</b>
Sulphurous acid Carbonic acid Phosphoric acid Nitrous acid	H <sub>2</sub> SO <sub>3</sub> H <sub>2</sub> CO <sub>3</sub> H <sub>3</sub> PO <sub>4</sub> HNO <sub>2</sub> Derived from sulphites Derived from carbonates Derived from phosphates Derived from nitrites	Ethanoic acid (CH <sub>3</sub> COOH)  Methanoic acid (HCOOH)	CITRIC ACID from lemons TARTARIC ACID from grapes ACETIC ACID from vinegar LACTIC ACID from sour milk Hydrochloric acid from digestive juices

<b>Whenever an acid is dissolved in water, it produces</b> HYDROGEN IONS	<b>Term given to the number of hydrogen ions produced by one molecule of an acid</b> BASICITY OF AN ACID	<b>Definition of basicity of an acid</b> BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> .
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# Guide Acids Bases And Salts

**DJ Losen**



## **Guide Acids Bases And Salts:**

This book delves into Guide Acids Bases And Salts. Guide Acids Bases And Salts is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Guide Acids Bases And Salts, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Guide Acids Bases And Salts
- Chapter 2: Essential Elements of Guide Acids Bases And Salts
- Chapter 3: Guide Acids Bases And Salts in Everyday Life
- Chapter 4: Guide Acids Bases And Salts in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Guide Acids Bases And Salts. The first chapter will explore what Guide Acids Bases And Salts is, why Guide Acids Bases And Salts is vital, and how to effectively learn about Guide Acids Bases And Salts.

3. In chapter 2, this book will delve into the foundational concepts of Guide Acids Bases And Salts. The second chapter will elucidate the essential principles that need to be understood to grasp Guide Acids Bases And Salts in its entirety.

4. In chapter 3, the author will examine the practical applications of Guide Acids Bases And Salts in daily life. The third chapter will showcase real-world examples of how Guide Acids Bases And Salts can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Guide Acids Bases And Salts in specific contexts. The fourth chapter will explore how Guide Acids Bases And Salts is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Guide Acids Bases And Salts. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Guide Acids Bases And Salts.

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