



FIG. 2 Example of a Gamma Log for the Hydrologic Observation Well KGS #1 Braun located near Hays, Kansas in the USA (in API units whereby SGR reflects the derived total gamma ray log (the sum of all the radiation contributions), and CGR reflects the computed gamma ray log (the sum of the potassium and thorium responses, leaving out the contribution from uranium).

3.2 Definitions of Terms Specific to This Standard:

3.2.1 *dead time, n*—the time after each pulse when a second pulse cannot be detected.

3.2.2 *dead time effect, n*—the inability to distinguish closely-spaced nuclear counts leads to a significant underestimation of gamma activity in high radiation environments and is known as the “dead time effect”.

3.2.3 *depth of investigation, n*—the radial distance from the measurement point to a point where the predominant measured response may be considered centered.

3.2.3.1 *Discussion*—The depth of investigation for borehole

logging is a radial distance from the borehole and is not to be confused with borehole depth or any depth measured from the surface.

3.2.4 *measurement resolution, n*—the minimum change in measured value that can be detected.

3.2.5 *vertical resolution, n*—the minimum thickness that can be separated into distinct units.

3.2.6 *volume of investigation, n*—the volume that contributes 90 % of the measured response.

3.2.6.1 *Discussion*—It is determined by a combination of theoretical and empirical modeling. The volume of investigation is non-spherical and has gradational boundaries.

4. Summary of Guide

4.1 This guide applies to borehole gamma logging and is to be used in conjunction with Guide D5753.

4.2 This guide briefly describes the significance and use, apparatus, calibration and standardization, procedures, and reports for conducting borehole gamma logging.

5. Significance and Use

5.1 An appropriately developed, documented, and executed guide is essential for the proper collection and application of gamma logs. This guide is to be used in conjunction with Guide D5753.

5.2 The benefits of its use include improving selection of gamma logging methods and equipment, gamma log quality and reliability, and usefulness of the gamma log data for subsequent display and interpretation.

5.3 This guide applies to commonly used gamma logging methods for geotechnical applications.

5.4 It is essential that personnel (see the Personnel section of Guide D5753) consult up-to-date textbooks and reports on the gamma technique, application, and interpretation methods.

6. Interferences

6.1 Most extraneous effects on gamma logs are caused by logging too fast, instrument problems, borehole conditions, and geologic conditions.

6.2 Logging too fast can significantly degrade the quality of gamma logs. Gamma counts originating at a given depth need to be averaged over a time interval such that the natural statistical variation in the rate of gamma photon emission is negligible (see Fig. 4).

6.3 Instrument problems include: a) electrical leakage of cable and grounding problems, b) degradation of detector efficiency attributed to loss of crystal transparency (fogging) or fractures or breaks in the crystal, and c) mechanical damage causing separation of crystal and photomultiplier tube.

6.4 Borehole conditions include: a) changes in borehole diameter (especially in the fluid-filled portion), b) casing type and number, c) radioactive elements in drilling fluid in the borehole, or in cement or slurry behind casing, d) steel casing or cement in the annulus around casing, and e) thickness of the annulus around casing.

Guide To Borehole Image Interpretation

**Charles Bertrand Schultz, T. M.
Stout, Society of Vertebrate
Paleontology**



Guide To Borehole Image Interpretation:

Guide To Borehole Image Interpretation Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Guide To Borehole Image Interpretation**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 °C and 64 °C respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus

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the purpose of the experiment is to understand what vapour pressure is and how it changes with temperature to understand the relationship between vapour pressure at a given temperature and boiling temperature at a given pressure to determine the vapour pressure of a pure liquid at various temperatures

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exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds to compare by following the trendline of the straight chain alkanes it could be concluded that as the molar mass of the straight chain alkanes increases their boiling points increase as well

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50 °C so in order to convert an experimental measurement taken in Calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that needs to be determined by physical properties such as boiling point melting point density solubility etc in this experiment I am to figure out the boiling point of a mystery substance in order to identify it chemical responsibility this experiment requires the handling of a Bunsen burner boiling

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there are a variety of methods by which a sample's boiling point can be determined including distillation reflux and by using a Thiele tube the most straightforward method uses a Thiele tube and has the advantage of using less than 0.5 mL of material