

Hydrolysis -

The Reaction of Ions with Water

Lab Report

Purpose

The purpose of this lab is to predict whether a substance is acidic, basic, or neutral and to use the universal indicator solution to test it.

Hypothesis

Sodium Acetate	$\text{NaCH}_3\text{COO}_{(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CH}_3\text{COO}^-_{(aq)}$ $\text{CH}_3\text{COO}^-_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{CH}_3\text{COOH}_{(aq)} + \text{OH}^-_{(aq)}$		Basic
Sodium Chloride	$\text{NaCl}_{(s)} \leftrightarrow \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$		Neutral
Ammonium Chloride	$\text{NH}_4\text{Cl}_{(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$		Acidic
Ammonium Sulfate	$(\text{NH}_4)_2\text{SO}_{4(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$	$K_a = 5.4 \times 10^{-10}$	Acidic
		$K_b = 8.3 \times 10^{-13}$	
Aluminum Chloride	$\text{AlCl}_{3(s)} \rightarrow \text{Al}^{3+}_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{Al}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Al}(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$		Acidic
Calcium Nitrate	$\text{Ca}(\text{NO}_3)_{2(s)} \rightarrow \text{Ca}^{2+}_{(aq)} + \text{NO}_{3(aq)}$		Neutral
Iron(III) Sulfate	$\text{Fe}_2(\text{SO}_4)_{3(s)} \rightarrow \text{Fe}^{3+}_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Fe}(\text{OH})^{2+}_{(aq)} + \text{H}^+_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HFe}^{2+}_{(aq)} + \text{OH}^-_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$	$K_a = 6.0 \times 10^{-3}$	Acidic
		$K_b = 1.7 \times 10^{-12}$	
		$K_b = 8.3 \times 10^{-13}$	
Sodium Carbonate	$\text{Na}_2\text{CO}_{3(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CO}_3^{2-}_{(aq)}$ $\text{CO}_3^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HCO}_3^-_{(aq)} + \text{OH}^-_{(aq)}$		Basic

Heath Chemistry Lab 20d Hydrolysis Answers

Jin-Ying Zhang



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The book delves into Heath Chemistry Lab 20d Hydrolysis Answers. Heath Chemistry Lab 20d Hydrolysis Answers is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Heath Chemistry Lab 20d Hydrolysis Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Heath Chemistry Lab 20d Hydrolysis Answers
- Chapter 2: Essential Elements of Heath Chemistry Lab 20d Hydrolysis Answers
- Chapter 3: Heath Chemistry Lab 20d Hydrolysis Answers in Everyday Life
- Chapter 4: Heath Chemistry Lab 20d Hydrolysis Answers in Specific Contexts
- Chapter 5: Conclusion

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- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Heath Chemistry Lab 20d Hydrolysis Answers.

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