



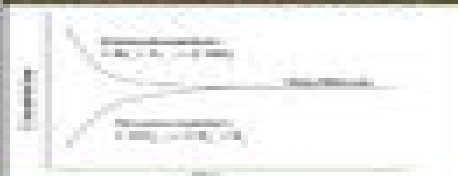
CHEMISTRY

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EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$.

What is Equilibrium? Equilibrium is a state of balance between two or more opposing processes. In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time. Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.	Equilibrium Constant The equilibrium constant, K, is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	
Equilibrium Constant, K_c K_c is the equilibrium constant based on concentrations. It is used for reactions in solution or in the gas phase. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant, K_p K_p is the equilibrium constant based on partial pressures. It is used for reactions in the gas phase. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	
Equilibrium Constant, K K is the equilibrium constant based on activities. It is used for reactions in solution or in the gas phase. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K = \frac{a_C^c a_D^d}{a_A^a a_B^b}$	Relationship between K_c and K_p The relationship between K_c and K_p is given by the equation: $K_p = K_c (RT)^{\Delta n}$ where R is the gas constant, T is the temperature in Kelvin, and Δn is the change in the number of moles of gas.	
Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Concentration	• Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products. • Increasing the concentration of a product will shift the equilibrium to the left, favoring the reactants.	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the concentration of N_2 will shift the equilibrium to the right, favoring the formation of NH_3.
Pressure	• Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas. • Decreasing the pressure will shift the equilibrium to the side with the most moles of gas.	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the pressure will shift the equilibrium to the right, favoring the formation of NH_3.
Temperature	• Increasing the temperature will shift the equilibrium to the side that is endothermic. • Decreasing the temperature will shift the equilibrium to the side that is exothermic.	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the temperature will shift the equilibrium to the left, favoring the reactants.

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