Chapter 9 Stoichiometry Chapter Review Answers

- 1. Mole Statios
 - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
 - They are formed from the coefficients found in a balanced chemical equation.
- 2. Mole Ratios as conversion factors
 - $a. \quad \frac{2 \bmod Ca}{1 \bmod O_2}, \frac{1 \bmod O_2}{2 \bmod Ca}, \frac{2 \bmod Ca}{2 \bmod Ca}, \frac{2 \bmod Ca}{2 \bmod Ca}, \frac{1 \bmod O_2}{2 \bmod CaO}, \frac{1 \bmod O_2}{1 \bmod O_2}, \frac{2 \bmod CaO}{1 \bmod O_2}, \\ Remember the ratios can be reduced as well 11 \bmod O_2$
 - $\mathbf{b.} = \frac{1 \bmod Hg}{2 \bmod HF}, \frac{2 \bmod HF}{1 \bmod Hg}, \frac{1 \bmod HgF_2}{1 \bmod Hg}, \frac{1 \bmod Hg}{1 \bmod HgF_2}, \frac{1 \bmod Hg}{1 \bmod Hg}, \frac{1 \bmod Hg}{1 \bmod Hg}, \frac{2 \bmod HF}{1 \bmod Hg}, \frac{2 \bmod HF}{1 \bmod HgF_2}, \frac{2 \bmod HF}{2 \bmod HF}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1$
- $3. \quad \frac{165.79 \, g \, Na_{2}CO_{3}}{1 \, mol \, Na_{2}CO_{3}}, \\ \frac{1}{105.79 \, g \, Na_{2}CO_{3}}, \\ \frac{74.10 \, g \, Co(0H)_{2}}{1 \, mol \, Co(0H)_{2}}, \\ \frac{1}{74.10 \, g \, Co(0H)_{2}}, \\ \frac{1}{74.10 \, g \, Co(0H)_{2}}, \\ \frac{1}{100.01 \, g \, NaOM}, \\ \frac{1}{100.01 \, g \, CoCO_{3}}, \\ \frac{1}{100.01 \, g \, CO$
- 4. Molar Mass
 - a. Definition: the mass of one mole of a substance
 - Molar mass is used as a conversion factor in stoiciometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen → Water
 - a. 5.0 mol H₂ b. 2.5 mol O₂
- 5. Ethane combustion reaction
 - a. 15.8 mol O₂
 b. 9.00 mol CO₂ & 13.5 mol H₂O
- 575 g Na & 886 g Cli
- 8. Inon production reaction
 - a. 75.1 mol CO
 b. 50.1 mol Fe & 75.1 mol CO₂
- 87,390 g CO or 8.739 × 10° g CO & 12,610 g H; or 1.261 × 10° g H;
- Lightning reaction of N₂ & O₂
 - a. 1,100 g NO₂ or 1.10 × 10³ g NO₃
- b. 720. g NO or 7.20 × 102 g NO

- 11. CO2 removal in spacecraft
 - a. 42.04 mol NaOH
- b. 21.02 mol Na₂CO₃ & 21.02 mol H₂O
- 12. Double replacement AgNO₃ & NaBr
 - a. 463 g NaBr
 b. 845 g AgBr
- 13. Soda-acid fire extinguisher
 - a. 3.058 mol NaHCO₃
- 3.058 mol CO₂, 1.529 mol Na₂SO₄, 3.058 mol H₂O
- 14. Sulfuric acid-sodium hydroxide neutralization reaction
 - a. H₂SO₀ + 2 NaOH → Na₂SO₀ + 2 H₂O
 - b. 37 g H₂SO₄
 - c. 53 g Na₁SO₁ & 14 g H₂O
- 15. Copper & silver nitrate reaction
 - a. 0.0104 mol Cu(NO₃)₂
 - b. 0.0104 mol Cu & 0.0209 mol AgNO₁
- 16. Aspirin production
 - a. 13.5 kg C₂H₂O₄
 - b. 7.66 kg C₄H₄O₃
 - 4.29 L HIC3H3O3
- 17. Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- 18. The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant fimits the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

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