

Name _____

Stoichiometry

21. How do you determine the amount of product formed in a reaction?

Use stoichiometry based on the limiting reactant.

22. What is the limiting reactant in $2SO_2 + O_2 \rightarrow 2SO_3$ if you start with 3 moles of SO_2 and 2 moles of O_2 ?

SO_2

23. Why is it important to balance a chemical equation for stoichiometry calculations?

To ensure the correct mole ratios are used.

24. How can you increase the percent yield of a reaction?

Optimize reaction conditions and minimize losses.

25. What information is needed to calculate theoretical yield?

Balanced chemical equation and amounts of reactants.

26. If the percent yield is 50% and the actual yield is 5 grams, what is the theoretical yield?

10 grams

27. What is the first step in stoichiometric calculations?

Balance the chemical equation.

28. In $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$, what is the mole ratio of O_2 to CO_2 ?

5:3

29. If you have 10 grams of $NaCl$ and the reaction has a percent yield of 75%, how much product is obtained?

7.5 grams

30. How do you identify the limiting reactant in a multi-step reaction?

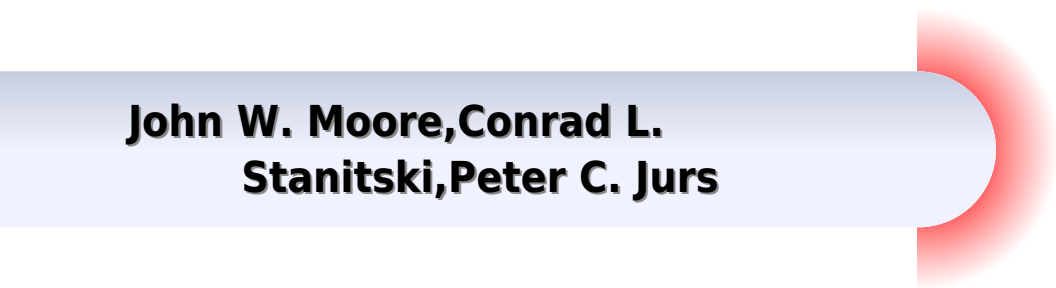
Determine the limiting reactant for each step and use it for calculations in subsequent steps.



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**John W. Moore, Conrad L.
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