

15. (Pre-lecture) Study the following reaction:
- a. Write out the **balanced equation**
b. Calculate ΔH_{rxn} **kJ/mol**
16. (Pre-lecture) This reaction is thermodynamically favored at 25°C. Is it also kinetically favored?
a. Explain why/why not.
b. How would you know?
c. How would you know?
d. How would you know?
17. (Pre-lecture) Calculate the equilibrium constant K_{eq} for the reaction of H_2O and CO_2 at 25°C. (Equilibrium constant K_{eq} for the reaction of H_2O and CO_2 is 1.0 at 25°C.)
a. The **equilibrium constant** is **1.0**
b. The **equilibrium constant** is **1.0**
c. The **equilibrium constant** is **1.0**
d. The **equilibrium constant** is **1.0**
18. (Pre-lecture) Based on the following data, calculate the ΔH_{rxn} for the reaction of H_2O and CO_2 at 25°C.
a. $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$
b. $\Delta H_{\text{f}}^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$
c. $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$
d. $\Delta H_{\text{f}}^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$

Reaction	Equilibrium constant K_{eq}
$\text{H}_2\text{O} + \text{CO}_2 \rightleftharpoons \text{H}_2\text{O} + \text{CO}_2$	1.0
$\text{H}_2\text{O} + \text{CO}_2 \rightleftharpoons \text{H}_2\text{O} + \text{CO}_2$	1.0
$\text{H}_2\text{O} + \text{CO}_2 \rightleftharpoons \text{H}_2\text{O} + \text{CO}_2$	1.0
$\text{H}_2\text{O} + \text{CO}_2 \rightleftharpoons \text{H}_2\text{O} + \text{CO}_2$	1.0

19. (Pre-lecture) The following data are for the reaction of H_2O and CO_2 at 25°C. Calculate the ΔH_{rxn} for the reaction of H_2O and CO_2 at 25°C. (Equilibrium constant K_{eq} for the reaction of H_2O and CO_2 is 1.0 at 25°C.)
a. $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$
b. $\Delta H_{\text{f}}^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$
c. $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$
d. $\Delta H_{\text{f}}^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$

$$\Delta H_{\text{rxn}} = \Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) + \Delta H_{\text{f}}^\circ(\text{CO}_2) - \Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) - \Delta H_{\text{f}}^\circ(\text{CO}_2) = -285.8 \text{ kJ/mol} - 393.5 \text{ kJ/mol} - (-285.8 \text{ kJ/mol}) - (-393.5 \text{ kJ/mol}) = 0 \text{ kJ/mol}$$

20. (Pre-lecture) Based on the following data, calculate the ΔH_{rxn} for the reaction of H_2O and CO_2 at 25°C.
a. $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$
b. $\Delta H_{\text{f}}^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$
c. $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$
d. $\Delta H_{\text{f}}^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$

$$\Delta H_{\text{rxn}} = \Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) + \Delta H_{\text{f}}^\circ(\text{CO}_2) - \Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) - \Delta H_{\text{f}}^\circ(\text{CO}_2) = -285.8 \text{ kJ/mol} - 393.5 \text{ kJ/mol} - (-285.8 \text{ kJ/mol}) - (-393.5 \text{ kJ/mol}) = 0 \text{ kJ/mol}$$

21. (Pre-lecture) Calculate the ΔH_{rxn} for the reaction of H_2O and CO_2 at 25°C. (Equilibrium constant K_{eq} for the reaction of H_2O and CO_2 is 1.0 at 25°C.)
a. $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$
b. $\Delta H_{\text{f}}^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$
c. $\Delta H_{\text{f}}^\circ(\text{H}_2\text{O}) = -285.8 \text{ kJ/mol}$
d. $\Delta H_{\text{f}}^\circ(\text{CO}_2) = -393.5 \text{ kJ/mol}$

The **equilibrium constant** is **1.0**
The **equilibrium constant** is **1.0**
The **equilibrium constant** is **1.0**
The **equilibrium constant** is **1.0**

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Ying Liu

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